

Introduction to Strong and Weak Acids and Bases PhET Lab (rvald 5/2011)*How does the strength of an acid or base affect conductivity? pH?***Introduction:**

When you test your pool's pH, what are those little vials or paper strips telling you? When you hear an acid called "strong" or "weak", what do those terms refer to? In aqueous solutions, compounds can exist as molecules (*undissociated*) or ions (*dissociated*). When an acid or a base exists in solution nearly completely as dissociated ions, we refer to that acid or base as *strong*. A *weak* acid or base will donate ions to the solution, but will remain primarily as undissociated molecules.

**Acid-Base Solutions****Notation:**

Acids are abbreviated HA, with the H representing the proton (H^+) the acid donates to the solution. The A is referred to as the acidic anion (A^-) that is left in solution as the proton is donated. $HA \rightleftharpoons H^+ + A^-$

Strong Bases are abbreviated MOH, with the OH representing the hydroxide ion (OH^-) the base donates to the solution. The M is cation (M^+) that is left in solution as the hydroxide is donated. $MOH \rightarrow M^+ + OH^-$.

**Autoionization:**

Even without any acid or base added a very small number of water molecules will form protons (H^+) and hydroxide ions (OH^-). The protons will then form **hydronium ions**, the acid ion.

Procedure: PhET Simulations → Play With Sims → Chemistry → Acid-Base Solutions → **Run Now!**

The concentration of the acids and bases used in the **Introduction** at 0.010 (10^{-2}) Molar.

- Begin with a **strong acid** and lower the pH probe into the beaker. What is the pH of this solution?
- Test this strong acid with both pH paper and the conductivity probe. What color does the pH indicator become? Is this strong acid an electrolyte? Does current travel through this solution?
- Repeat the above tests with the weak acid, the strong base, and the weak base, and water. Collect your observations in the table below:

	Strong Acid	Weak Acid	Strong Base	Weak Base	Water
pH meter read (value)	2.00	4.50	12.00	9.50	7.00
pH paper (color)	2 red	4 orange	11 blue	9 green	7 yellow
Conductivity (bright/dim/none)	Bright	dim	10 bright	dim	none
Exists as Mostly (ions/molecules)	ions	ions	molecules	molecules	ions

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Table of Contents Answer Key To Phet Densitysolution Answer

1. Understanding the eBook Answer Key To Phet Densitysolution Answer
 - The Rise of Digital Reading Answer Key To Phet Densitysolution Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key To Phet Densitysolution Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key To Phet Densitysolution Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key To Phet Densitysolution Answer
 - Personalized Recommendations
 - Answer Key To Phet Densitysolution Answer User Reviews and Ratings
 - Answer Key To Phet Densitysolution Answer and Bestseller Lists

5. Accessing Answer Key To Phet Densitysolution Answer Free and Paid eBooks
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 - Answer Key To Phet Densitysolution Answer eBook Subscription Services
 - Answer Key To Phet Densitysolution Answer Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Answer Key To Phet Densitysolution Answer Compatibility with Devices
 - Answer Key To Phet Densitysolution Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key To Phet Densitysolution Answer
 - Highlighting and Note-Taking Answer Key To Phet Densitysolution Answer
 - Interactive Elements Answer Key To Phet Densitysolution Answer
8. Staying Engaged with Answer Key To Phet Densitysolution Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key To Phet Densitysolution Answer
9. Balancing eBooks and Physical Books Answer Key To Phet Densitysolution Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key To Phet Densitysolution Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key To Phet Densitysolution Answer
 - Setting Reading Goals Answer Key To Phet Densitysolution Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key To Phet Densitysolution Answer
 - Fact-Checking eBook Content of Answer Key To Phet Densitysolution Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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