Same _____

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- a. Write the balanced chemical equation
- b. Identify the given (with units) and what you want to find (with units)
- c. Show set up with units. Check sig figs, give final answer with units and label.
- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?

Equation:	2 H ₂ O _(i) →	2 H2(c) +	1 O2 (a)	
Before	2 mol	0 mod	0 mol	
Change	-2 mol	+2 mol		
After	0 mol	2 mol		

$$36 \text{ g } H_2Ox \frac{1 \text{ mol } H_2O}{18.0 \text{ g } H_2O} = 2.0 \text{ mol } H_2O \times \frac{2 \text{ mol } H_2}{2 \text{ mol } H_2O} = 2 \text{ mol } H_2 \times \frac{2.0 \text{ g } H_2}{1 \text{ mol } H_2} = 4.0 \text{ g } H_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?

Equation:	2 Na ₍₀₎ +	$-$ L $Cl_{2(g)} \rightarrow$		100
Before	9.94 mol	xs mol	0 mod	
Change	-9.94 mod		+9.94 mol	
After	0 mod		9.94 mod	

$$581 \ g \ NaClx \\ \frac{1 \ mol \ NaCl}{58.5 g \ NaCl} = 9.93 \ mol \ NaCl \\ \times \\ \frac{2 \ mol \ Na}{2 \ mol \ NaCl} = 9.93 \ mol \ Na \\ \times \\ \frac{23.0 \ g \ Na}{1 \ mol \ Na} = 228 \ g \ Na \\ \frac{1 \ mol \ Na}{2 \ mol \ Na} = 228 \ g \ Na \\ \times \\ \frac{1 \ mol \ Na}{2 \ mol \ N$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose, C6H12O6, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?

Equation:	1 CeH ₁₂ O _{6(e)} +	6 O₂(g) →	6 CO _{2(g)}	+ 6 H ₂ O _(g)
Before	1.0 mod	xs mol	0 mod	0 mol
Change	-1.0 mol	-1.0 mol		- 30.00.00
After	0 mol	xs mol		

$$180.0 \text{ g C}_{a}H_{12}O_{a}x \frac{1 \text{ mol C}_{a}H_{12}O_{a}}{180 \text{ g C}_{a}H_{12}O_{a}} = 1.00 \text{ mol C}_{a}H_{12}O_{a} \times \frac{6 \text{ mol } O_{2}}{1 \text{ mol C}_{a}H_{12}O_{a}} = 6.00 \text{ mol } O_{2}$$

6.0 mol
$$O_2 x \frac{32.0 \text{ g O}_2}{1 \text{ mol O}_2} = 192.0 \text{ g } O_2$$

Answer Key Chemical Equations And Stoichiometry

Purshotam Kaushik & Kirti Kaushik

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disciplines such as biotechnology polymer technology petrochemical engineering electrochemical engineering environmental engineering and safety engineering the chief objective of the book is to prepare students to make analysis of chemical processes through calculations and to develop systematic problem solving skills in them The text presents the fundamentals of chemical engineering operations and processes in a simple style that helps the students to gain a thorough understanding of chemical process calculations The book deals with the principles of stoichiometry to formulate and solve material and energy balance problems in processes with and without chemical reactions With the help of examples the book explains the construction and use of reference substance plots equilibrium diagrams psychrometric charts steam tables and enthalpy composition diagrams It also elaborates on thermophysics and thermochemistry to acquaint the students with the thermodynamic principles of energy balance calculations The book is supplemented with Solutions Manual for instructors containing detailed solutions of all chapter end unsolved problems NEW TO THE SECOND EDITION Incorporates a new chapter on Bypass Recycle and Purge Operations Comprises updations in some sections and presents new sections on Future Avenues and Opportunities in Chemical Engineering Processes in Biological and Energy Systems Contains several new worked out examples in the chapter on Material Balance with Chemical Reaction Includes GATE questions with answers up to the year 2016 in Objective type questions KEY FEATURES SI units are used throughout the book All basic chemical engineering operations and processes are introduced and different types of problems are illustrated with worked out examples Stoichiometric principles are extended to solve problems related to bioprocessing environmental engineering etc Exercise problems more than 810 are organised according to the difficulty level and all are provided with answers

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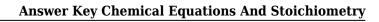
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