

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ for a reaction is 2074° whereas a 4.000 g of substance reacts with 10.00 g of O_2 to produce 14.00 g of substance. What is the $\Delta H_{\text{f}}^\circ$ for the substance?
- _____ A substance has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If a 1.0 g sample of substance is heated from 25°C to 100°C , how much heat is absorbed? What is the $\Delta H_{\text{f}}^\circ$ for the substance?
- The product is a molecular formula of $\text{C}_2\text{H}_4\text{O}$, and the substance has a molar mass of 44.05 g/mol .
 - What is the value of $\Delta H_{\text{f}}^\circ$ for this reaction?

Answers For Reaction Energy Section 1

G Thomas



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