Name: Date: Period:

## POGIL- Stoichiometry

### He got Mol-aria

What happened to

Avogadro when he got bit by 6.02 x

10 " mosquitoes?

How do chemists use balanced chemical equations?

### Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called <u>Stoichiometry</u> (stoi-key-ah-meh-tree) Another way of looking at it is using the <u>mole ratio</u> from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A <u>mole ratio</u> is a conversion factor <u>derived from the coefficients of a balanced chemical equation</u> interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: How do I convert between different chemical species in a given reaction?

### MODEL 1: Mole-to-Mole Stoichiometry (1-step)

# Example: mole-to-mole Question: Given the following equation: $1 N_2(g) + 3 H_2(g) \rightarrow 2 NH_3(g)$ How many moles of $H_2$ are needed to produce 34.8 moles of $NH_3$ ? Mole ratio Answer: $34.8 \text{ mol NH}_3 \mid 3 \text{ mol H}_2 = 52.2 \text{mol NH}_2$ OR $34.8 \text{ mol NH}_3 \times 3 \text{ mol H}_2 = 52.2 \text{mol NH}_3$

### Key Questions:

- 1. What was the given in the example above?
- Where did the <sup>1 moln's</sup>/<sub>2 moln's</sub>, come from?
- Double check to see if 52.2 moles of H<sub>2</sub> is the correct answer. <u>SHOW</u> the math below.

# **Answer For Basic Stoichiometry Pogil Activity**

Minu Gupta Bhowon, Sabina Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami

### **Answer For Basic Stoichiometry Pogil Activity:**

Chemistry Education in the ICT Age Minu Gupta Bhowon, Sabina Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami, 2009-07-21 th th The 20 International Conference on Chemical Education 20 ICCE which had rd th Chemistry in the ICT Age as the theme was held from 3 to 8 August 2008 at Le M ridien Hotel Pointe aux Piments in Mauritius With more than 200 participants from 40 countries the conference featured 140 oral and 50 poster presentations th Participants of the 20 ICCE were invited to submit full papers and the latter were subjected to peer review The selected accepted papers are collected in this book of proceedings This book of proceedings encloses 39 presentations covering topics ranging from fundamental to applied chemistry such as Arts and Chemistry Education Biochemistry and Biotechnology Chemical Education for Development Chemistry at Secondary Level Chemistry at Tertiary Level Chemistry Teacher Education Chemistry and Society Chemistry Olympiad Context Oriented Chemistry ICT and Chemistry Education Green Chemistry Micro Scale Chemistry Modern Technologies in Chemistry Education Network for Chemistry and Chemical Engineering Education Public Understanding of Chemistry Research in Chemistry Education and Science Education at Elementary Level We would like to thank those who submitted the full papers and the reviewers for their timely help in assessing the papers for publication th We would also like to pay a special tribute to all the sponsors of the 20 ICCE and in particular the Tertiary Education Commission http tec intnet mu and the Organisation for the Prohibition of Chemical Weapons http www opcw org for kindly agreeing to fund the publication of these proceedings Process Oriented Guided Inquiry Learning (POGIL) Richard Samuel Moog, 2008 POGIL is a student centered group learning pedagogy based on current learning theory This volume describes POGIL s theoretical basis its implementations in diverse environments and evaluation of student outcomes

Calculating in Chemistry Cosmas C Mbuzi,2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to covert various quantities to moles or amount of substance This particularly useful to a beginner Chapter 2 describes how to calculate reacting reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various

circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry

A Stoichiometry Unit David Callaghan, 2004

Formulation and Stoichiometry Emil J. Margolis,1968 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in con trasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction

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