

Reviewing Content

42. The solvent is the substance in which the solute is dissolved.
43. Random collisions of the solvent molecules with the solute particles provide enough force to overcome gravity.
44. **solubility**: the amount of a substance that dissolves in a given quantity of solvent at specified conditions of temperature and pressure to produce a saturated solution.
saturated solution: a solution containing the maximum amount of solute for a given amount of solvent at a constant temperature and pressure. **unsaturated solution**: a solution that contains less solute than a saturated solution at a given temperature and pressure. **miscible**: describes liquids that dissolve in each other. **immiscible**: describes liquids that are insoluble in each other.
45. Particles of solute crystallize.
46. No; if there were undissolved solute, the excess solute would come out of a supersaturated solution.
47. 5.55×10^2 g AgNO_3
48. Solubility increases with pressure.
49. a. 1.6×10^{-2} g/L
b. 4.7×10^{-2} g/L
50. *Dilute* and *concentrated* are relative terms and are not quantitative. Molarity provides the exact number of moles of solute per liter of solution.
51. Molarity is the number of moles of solute dissolved in one liter of solution.
a. 1.3M KCl
b. 3.3×10^{-1} M MgCl_2
52. 2.00×10^3 mL
53. a. 5.0×10^{-3} mol NaCl, 29 g NaCl
b. 1.0 mol KNO_3 , 1.0×10^{-2} g KNO_3
c. 2.5×10^3 mol CaCl_2 , 2.8 g CaCl_2
54. a. 2.3×10^1 g NaCl
b. 2.0 g MgCl_2
55. a. 16% (v/v) ethanol
b. 63.6% (v/v) isopropyl alcohol
56. Colligative properties are properties of a solution that depend only on the number of solute particles; boiling-point elevation, freezing-point depression, and vapor-pressure lowering. Boiling points are elevated because shells of solvent form around solute particles, reducing the amount of solvent molecules that have sufficient energy to escape the solution; relative to the pure solvent, the amount of energy required to cause vaporization or boiling increases. Solutes disrupt the ordering of the solvent structure, so more kinetic energy must be withdrawn from a solution for it to solidify. This lowers the freezing point of the solution.
57. a. sea water
b. 1.5M KNO_3
c. 0.100M MgCl_2
58. The effective molality of the $\text{Ca}(\text{NO}_3)_2$ solution is 3m. The effective molality of the NaNO_3 solution is 2m.
59. When vapor pressure is lowered relative to pure solvent, more energy must be supplied to reach the boiling point; thus the boiling point is increased relative to pure solvent.
60. The salt lowers the freezing point of the ice-water cooling mixture.
61. 1M solution: 1 mol of solute in 1 L of solution; 1m solution: 1 mol of solute in 1000 g of solvent
62. Add 27.0 g H_2O to 32.0 g CH_3OH .
63. a. 100.26°C
b. 101.54°C
64. a. -4.46°C
b. -2.2°C
65. a. -1.1°C
b. -0.74°C
c. -1.5°C

Understanding Concepts

66. a. The freezing-point depression is twice as great for solute B; solute B must provide twice as many particles in solution.

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Table of Contents Answer Key For Chemistry Ch 16 Assessment

1. Understanding the eBook Answer Key For Chemistry Ch 16 Assessment
 - The Rise of Digital Reading Answer Key For Chemistry Ch 16 Assessment
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Chemistry Ch 16 Assessment
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Chemistry Ch 16 Assessment
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Chemistry Ch 16 Assessment
 - Personalized Recommendations
 - Answer Key For Chemistry Ch 16 Assessment User Reviews and Ratings
 - Answer Key For Chemistry Ch 16 Assessment and Bestseller Lists
5. Accessing Answer Key For Chemistry Ch 16 Assessment Free and Paid eBooks
 - Answer Key For Chemistry Ch 16 Assessment Public Domain eBooks
 - Answer Key For Chemistry Ch 16 Assessment eBook Subscription Services
 - Answer Key For Chemistry Ch 16 Assessment Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Answer Key For Chemistry Ch 16 Assessment Compatibility with Devices
 - Answer Key For Chemistry Ch 16 Assessment Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Answer Key For Chemistry Ch 16 Assessment
 - Interactive Elements Answer Key For Chemistry Ch 16 Assessment
8. Staying Engaged with Answer Key For Chemistry Ch 16 Assessment
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key For Chemistry Ch 16 Assessment
9. Balancing eBooks and Physical Books Answer Key For Chemistry Ch 16 Assessment
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For Chemistry Ch 16 Assessment
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answer Key For Chemistry Ch 16 Assessment
 - Setting Reading Goals Answer Key For Chemistry Ch 16 Assessment
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Key For Chemistry Ch 16 Assessment
 - Fact-Checking eBook Content of Answer Key For Chemistry Ch 16 Assessment
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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