

Chemistry Worksheets Class 11 on Chapter 4 Chemical Bonding and Molecular Structure - Set 3

Q-1: Which of the following possesses an expanded octet?

- a) IF_7
- b) SF_6
- c) BrF_5
- d) All of the above

Q-2: Strongly electronegative element B contrasts with strongly electropositive element A. Both are divalent. The resulting compound would be

- a) $\text{A} \rightarrow \text{B}$
- b) $\text{A}^{2+}\text{B}^{2-}$
- c) $\text{A}^{2-}\text{B}^{2+}$
- d) $\text{A}-\text{B}$

Q-3: Which of the following compounds shows the highest lattice energy?

- a) CsCl
- b) KCl
- c) RbCl
- d) NaCl

Q-4: Which one of the following molecules is formed by p-p overlapping?

- a) I_2
- b) HBr
- c) H_2S
- d) PH_3

Q-5: Which of the following H-bonds has the greatest impact on a molecule's physical characteristics?

- a) Intramolecular H-bond
- b) Intermolecular H-bond
- c) Both of the above
- d) All H-bonds

Q-6: State the crucial conditions that must be met for a molecule to undergo hybridisation.

Q-7: Describe why CH_4 has a tetrahedral geometry rather than a square planar geometry with a carbon atom in the centre and four H atoms at each corner.

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Wolfgang Guggemos



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