

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a sample solution containing chloride ions. The sample solution is a mixture of sodium chloride and sodium nitrate. The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride in the original sample. The experiment is performed in two parts. In the first part, the sample is weighed and dissolved. In the second part, the sample is treated with silver nitrate and the precipitate is filtered, dried, and weighed.

The sample is weighed and dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride in the original sample.

Procedure

The procedure for this experiment is as follows: 1. Weigh a sample of the solution. 2. Dissolve the sample in water. 3. Add silver nitrate to the solution. 4. Filter the precipitate. 5. Dry the precipitate. 6. Weigh the precipitate.

Data Analysis

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride in sample	0.350 g	0.350 g

The mass of chloride in the sample is calculated as follows: $m_{\text{Cl}^-} = m_{\text{precipitate}} \times \frac{m_{\text{Cl}^-}}{m_{\text{AgCl}}}$

For Sample 1: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45}{143.32} = 0.350 \text{ g}$

For Sample 2: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45}{143.32} = 0.350 \text{ g}$

Analysis Of Chloride Experiment 7 Lab Report

M Mosston



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