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Analysis, Synthesis, and Design of Chemical Processes

Third Edition

Prentice Hall International Series
in the Physical and Chemical
Engineering Sciences



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web we utilize the equation that relate amount remaining initial mass and number of half lives $n = \frac{t}{t_{1/2}} \times \log_2 \frac{N_0}{N}$ 125 g 5 what is the half life of an isotope that is 80% remained after 16 days solution remaining 80 100 therefore $n = \frac{t}{t_{1/2}} \times \log_2 \frac{N_0}{N}$ now using the half life equation in example 1 we have

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