

Question 1

1(a)(iii) Difference: With powdered CaCO_3 the graph has a steeper slope at the start of the reaction than with marble chips because the rate of reaction is faster due to increased surface area of the solid.
Similarity: The total volume of CO_2 produced is the same in both cases since the same mass (and therefore number of moles) of CaCO_3 was used in both experiments.



1(a)(v) Moles of CaCO_3 used = $(1.8/100)$ mol. = 0.018 mol.
From the balanced equation for the reaction, 1 mole CaCO_3 produces 1 mol CO_2 .
Hence volume of gas produced at R.T.P. = (0.018×24000) cm^3 = 432 cm^3 .

1(a)(vi) There were impurities in the CaCO_3 used.

1(a)(vii) Concentration of the acid used and temperature at which the reaction was carried out.

1(b)(i) $\text{Fe}^{3+}_{(\text{aq})}$ present.

1(b)(ii) Magnesium ribbon decreases in size/dissolves.
Solution becomes pale green.



1(b)(iv) SO_4^{2-} , SO_3^{2-} and CO_3^{2-} absent.

1(b)(v) White precipitate, which is soluble in aqueous ammonia, is formed.

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Table of Contents Answers For January 2014 Chemistry

1. Understanding the eBook Answers For January 2014 Chemistry
 - The Rise of Digital Reading Answers For January 2014 Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For January 2014 Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For January 2014 Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For January 2014 Chemistry
 - Personalized Recommendations
 - Answers For January 2014 Chemistry User Reviews and Ratings
 - Answers For January 2014 Chemistry and Bestseller Lists
5. Accessing Answers For January 2014 Chemistry Free and Paid eBooks
 - Answers For January 2014 Chemistry Public Domain eBooks
 - Answers For January 2014 Chemistry eBook Subscription Services
 - Answers For January 2014 Chemistry Budget-Friendly Options
6. Navigating Answers For January 2014 Chemistry eBook Formats

- ePub, PDF, MOBI, and More
- Answers For January 2014 Chemistry Compatibility with Devices
- Answers For January 2014 Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For January 2014 Chemistry
 - Highlighting and Note-Taking Answers For January 2014 Chemistry
 - Interactive Elements Answers For January 2014 Chemistry
- 8. Staying Engaged with Answers For January 2014 Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For January 2014 Chemistry
- 9. Balancing eBooks and Physical Books Answers For January 2014 Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For January 2014 Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For January 2014 Chemistry
 - Setting Reading Goals Answers For January 2014 Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For January 2014 Chemistry
 - Fact-Checking eBook Content of Answers For January 2014 Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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