# STOICHIOMETRY

 Calculate the number of moles of NaOH that are needed to react with 500 g of H₂SO₄ according to the following equation:

$$H_2SO_4 + 2 NaOH \rightarrow Na_2SO_4 + 2 H_2O$$

2. Calculate the mass of  $NH_3$  that can be produced from the reaction of  $125\,g$  of  $NCl_3$  according to the following equation:

3. Identify the limiting reactant and determine the mass of  $CO_2$  that can be produced from the reaction of 25 g of  $C_3H_8$  with 75 g of  $O_2$  according to the following equation:

$$C_3H_8 + 5O_2 \rightarrow 3 CO_2 + 4 H_2O$$

4. How many grams of  $SO_2$  are produced when 152 g of  $CS_2$  reacts with 48.0 g of  $O_2$  according to the following equation:

$$CS_2 + 3O_2 \rightarrow CO_2 + 2SO_2$$

# **Answers For The Chemistry Lab Of Stoichiometry**

**Dennis D. Staley, Edward L. Waterman** 

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