

Quick Lab

Purpose To use a graph to identify period and group trends

Materials

- graph paper
- pencil



Periodic Trends in Ionic Radii

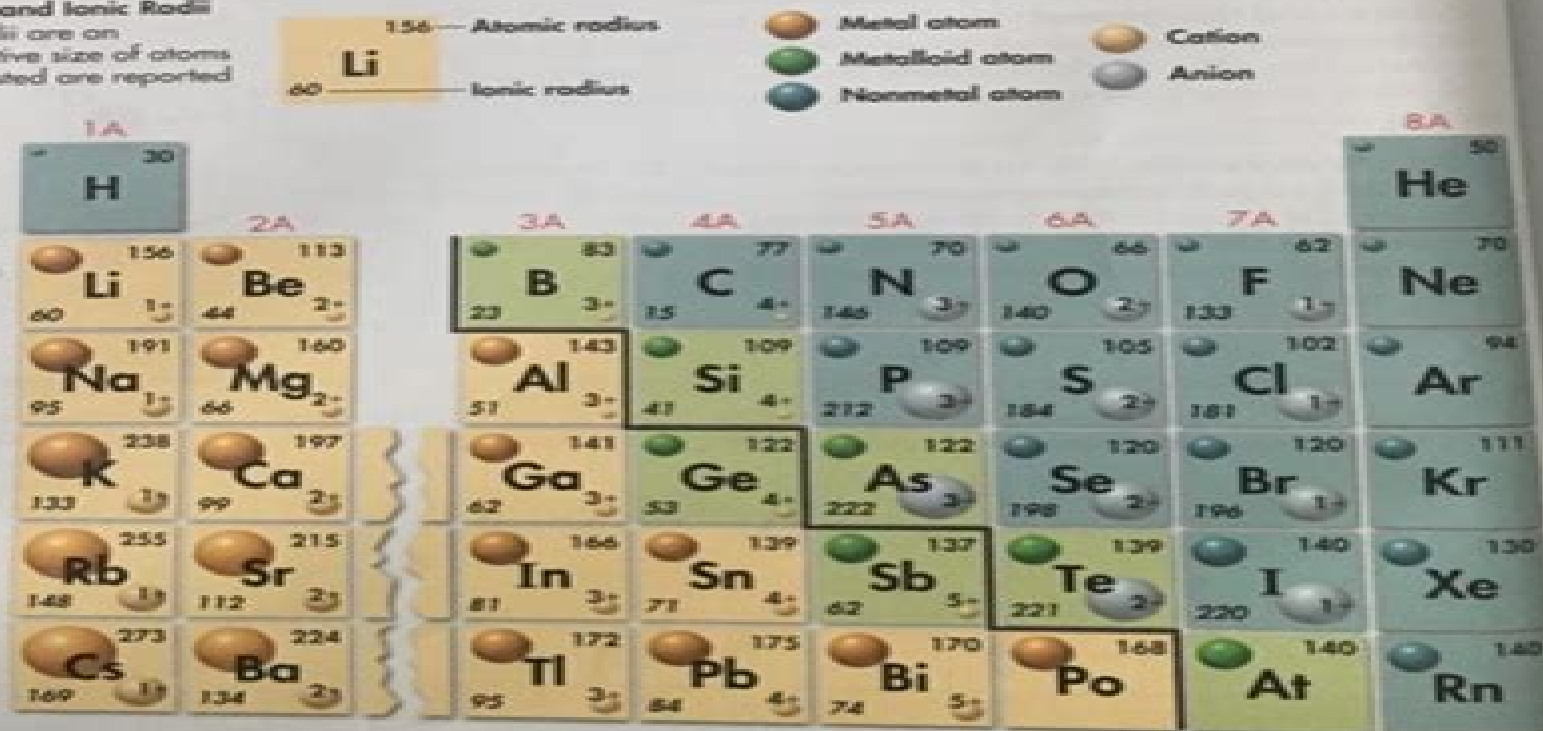
Procedure

Use the data presented in Figure 6.23 to plot ionic radius versus atomic number.

Analyze and Conclude

- 1. Compare** How does the size change when an atom forms a cation and when an atom forms an anion?
- 2. Describe** How do the ionic radii vary within a group of metals? How do they vary within a group of nonmetals?
- 3. Describe** What is the shape of a portion of the graph that corresponds to one period?
- 4. Compare and Contrast** Is the trend across a period similar or different for Periods 2, 3, 4, and 5?
- 5. Explain** Propose explanations for the trends you have described for ionic radii within groups and across periods.

Figure 6.23 Atomic and Ionic Radii
Atomic and ionic radii are an indication of the relative size of atoms and ions. The data listed are reported in picometers (pm).



Answers Quick Lab Periodic Trends In Ionic Radii

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



Answers Quick Lab Periodic Trends In Ionic Radii:

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The book delves into Answers Quick Lab Periodic Trends In Ionic Radii. Answers Quick Lab Periodic Trends In Ionic Radii is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Answers Quick Lab Periodic Trends In Ionic Radii, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 1: Introduction to Answers Quick Lab Periodic Trends In Ionic Radii
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 - Chapter 3: Answers Quick Lab Periodic Trends In Ionic Radii in Everyday Life
 - Chapter 4: Answers Quick Lab Periodic Trends In Ionic Radii in Specific Contexts
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