Stoichiometry Lab: Vinegar and Baking Soda Do Not Write On This Sheet

<u>Purpose</u>: To predict the amount of Carbon Dioxide gas that should be produced in a chemical reaction; then calculate the % yield.

CH, COOH + NaHCO; ---> NaCH, COO + H,O + CO,

Materials: Baking Soda (NaHCO₁), Vinegar (CH₃COOH), and 2 plastic cups, scale.

<u>Procedure</u>: Find and record the mass of cup A. With cup A still on the scale, add approximately 10.0 g of baking soda to the cup. The mass does not have to be exact, as long as you carefully record your results.

Place cup B on the scale, weigh and record approximately 50.0 g of vinegar. The mass does not have to be exact, as long as you carefully record your results.

Slowly add vinegar to cup A until the reaction has stopped. DO NOT add all of the vinegar, just enough to complete the reaction. Reweigh and record both cup A and B. Calculate the mass of CO₂ that escaped.

Data:

- a. Mass of Cup A _____ g
- b. Mass of Cup A and baking soda g
- c. Calculate mass of baking soda (b a) _____ g
- d. Mass of Cup B with vinegar g
- e. Mass of Cup B after reaction g
- f. Calculate mass of vinegar poured into Cup A (d e) g
- g. Mass of Cup A after reaction _____g
- b. Calculate mass of product after reaction (g a) ______g
- i. Calculate baking soda + vinegar (c + f) _____g
- j. Calculate mass of CO₂ lost (i h) _____g

Analysis Of Baking Soda Lab Answers

Kathleen Armour

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