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ppm Answer- $(\text{grams analyte}/\text{grams sample}) \times 10^6$

Molarity Answer- moles analyte/liter of solution

Volume Percent Answer- $(\text{volume solute}/\text{volume solution}) \times 100$

Volume ppm Answer- $(\text{volume solute}/\text{volume solution}) \times 10^6$

kilo- Answer- 10^3

deci- Answer- 10^{-1}

centi- Answer- 10^{-2}

milli- Answer- 10^{-3}

micro- Answer- 10^{-6}

nano- Answer- 10^{-9}

pico- Answer- 10^{-12}

femto- Answer- 10^{-15}

weight percent Answer- $(\text{grams analyte}/\text{grams sample}) \times 100$

ppt Answer- $(\text{grams analyte}/\text{grams sample}) \times 10^3$

ppt simplified Answer- gram analyte/liter solution

Analytical Chemistry Exam

Frederick Pearson Treadwell



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Sciences Corps He is an American Chemical Society member Dr Frame received the B A degree in chemistry from Harvard College Cambridge Massachusetts and the Ph D degree in analytical chemistry from Rutgers University New Brunswick New Jersey

A Text-book of Elementary Analytical Chemistry, Qualitative and Volumetric John Harper Long, 1898

TRAC: Trends in Analytical Chemistry U A Th Brinkman, J. G. Dorsey, J R Durig, 2013-09-24

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