

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{mole of sodium chloride} = 25.0 \text{ mL NaCl} \times \frac{1.0 \text{ g/mL}}{58.44 \text{ g/mol}} = 0.428 \text{ mol NaCl}$$

$$\text{mole of calcium chloride} = 25.0 \text{ mL CaCl}_2 \times \frac{1.0 \text{ g/mL}}{110.98 \text{ g/mol}} = 0.225 \text{ mol CaCl}_2$$

$$\text{mole of sodium hydroxide produced} = 5.00 \text{ g NaOH} \times \frac{1 \text{ mol}}{40.00 \text{ g/mol}} = 0.125 \text{ mol NaOH}$$

$$\text{mole of product that would be produced} = 0.125 \text{ mol NaOH} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.125 \text{ mol NaOH}$$

2. Reaction equation: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Reaction equation:

$$\text{mole of Na}_2\text{CO}_3 = 5.00 \text{ g Na}_2\text{CO}_3 \times \frac{1 \text{ mol}}{105.99 \text{ g/mol}} = 0.0472 \text{ mol Na}_2\text{CO}_3$$

$$\text{mole of Na}_2\text{CO}_3 = 0.0472 \text{ mol Na}_2\text{CO}_3$$

$$\text{mole of Na}_2\text{CO}_3 \text{ that would be produced} = 0.0472 \text{ mol Na}_2\text{CO}_3 \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.0472 \text{ mol Na}_2\text{CO}_3$$

Answers For Lab 6b Heath Chemistry

Rosina Ehmann



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