

Name _____ answer key _____ Date _____

Stoichiometry: Mass – Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₃.

$$\frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3} \times \frac{1 \text{ mole FeCl}_3}{9 \text{ FeCl}_3} \times \frac{\text{mole NaCl}}{\text{mole FeCl}_3}$$

Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

$$\frac{9 \text{ Na}_2\text{SO}_4}{1 \text{ mole Na}_2\text{SO}_4} \times \frac{1 \text{ mole Na}_2\text{SO}_4}{9 \text{ Na}_2\text{SO}_4} \times \frac{\text{mole Fe}_2(\text{SO}_4)_3}{\text{mole Na}_2\text{SO}_4} \times \frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3}$$

Calculate the amount in grams of FeCl₃ produced with 95g of Fe (SO₄)₃

$$\frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3} \times \frac{1 \text{ mole Fe}_2(\text{SO}_4)_3}{9 \text{ Fe}_2(\text{SO}_4)_3} \times \frac{\text{mole FeCl}_3}{\text{mole Fe}(\text{SO}_4)_3} \times \frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3}$$

Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl

$$\frac{9 \text{ NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole NaCl}}{9 \text{ NaCl}} \times \frac{\text{mole Fe}(\text{SO}_4)_3}{\text{mole NaCl}}$$

Based on the following chemical equation, please answer the following questions!



$$\frac{9 \text{ Ca}(\text{HCO}_3)_2}{1 \text{ mole Ca}(\text{HCO}_3)_2} \times \frac{1 \text{ mole Ca}(\text{HCO}_3)_2}{9 \text{ Ca}(\text{HCO}_3)_2} \times \frac{\text{mole CaCO}_3}{\text{mole Ca}(\text{HCO}_3)_2}$$

Answers For Study Guide Mass Moles

Susan McMurry



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Table of Contents Answers For Study Guide Mass Moles

1. Understanding the eBook Answers For Study Guide Mass Moles
 - The Rise of Digital Reading Answers For Study Guide Mass Moles
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Study Guide Mass Moles
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Study Guide Mass Moles
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Study Guide Mass Moles
 - Personalized Recommendations
 - Answers For Study Guide Mass Moles User Reviews and Ratings
 - Answers For Study Guide Mass Moles and Bestseller Lists
5. Accessing Answers For Study Guide Mass Moles Free and Paid eBooks
 - Answers For Study Guide Mass Moles Public Domain eBooks
 - Answers For Study Guide Mass Moles eBook Subscription Services
 - Answers For Study Guide Mass Moles Budget-Friendly Options
6. Navigating Answers For Study Guide Mass Moles eBook Formats

- ePub, PDF, MOBI, and More
- Answers For Study Guide Mass Moles Compatibility with Devices
- Answers For Study Guide Mass Moles Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Study Guide Mass Moles
 - Highlighting and Note-Taking Answers For Study Guide Mass Moles
 - Interactive Elements Answers For Study Guide Mass Moles
- 8. Staying Engaged with Answers For Study Guide Mass Moles
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Study Guide Mass Moles
- 9. Balancing eBooks and Physical Books Answers For Study Guide Mass Moles
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Study Guide Mass Moles
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Study Guide Mass Moles
 - Setting Reading Goals Answers For Study Guide Mass Moles
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers For Study Guide Mass Moles
 - Fact-Checking eBook Content of Answers For Study Guide Mass Moles
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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