

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Sub-levels have different amounts of energy. For example, orbitals closer to the nucleus have more energy than orbitals further from the nucleus. The following orbitals have different levels of energy:



For each pair, write sublevel(s) for which orbitals are included. Then draw orbital(s) representing an orbital.



Now list 4 orbitals which appear in the set of pairs. Write down their approximate energy levels. For example, the 3d orbitals have higher energy than s orbitals and the 4f orbitals have a higher energy than d orbitals.

When electrons occupy orbitals, they try to have the lowest amount of energy possible. (This is called the **lowest-energy principle**.) An electron will occupy 1 orbital before the 1 orbital is filled and no electron will occupy an orbital until it is completely filled. Additionally, no electron will be sub-filled. (This is called the **Mutual Exclusion Principle**.) When two electrons occupy the same orbital they must spin in opposite directions – one clockwise and the other anticlockwise.

Critical Thinking Questions

- a) How many electrons would it take to fill a sub-level before beginning to be sub-filled? **Sub-levels become filled if the orbitals are filled before electrons reach orbitals in the 2p orbitals.**
b) What is the first element before enough electrons are present to fill 1 orbital? **After you provide this 1.**
The 1st 2p electron goes into the 1 orbital. The element is 1, hydrogen 1 (H).
- a) How many electrons would it take to fill a sub-level before beginning to be sub-filled? **Sub-levels become filled if the orbitals are filled before electrons reach orbitals in the 2p orbitals.**
b) What is the first element before enough electrons are beginning to fill orbitals in the 2p orbitals? **The 2nd 2p electron goes into the 2 orbitals. The element is 2, helium 2 (He).**

Information: Hund's Rule

Electrons prefer to "spin" or "upspin". Orbitals closer to the nucleus have a higher potential. Electrons closer to the nucleus have a lower potential. Electrons closer to the nucleus have a lower potential. Electrons closer to the nucleus have a lower potential.



If an orbital is represented as a box, an orbital closer to the nucleus will get electrons' orbital diagram. The first orbital is the one closest to the nucleus. The second orbital is the one closest to the nucleus.

Choice A or Choice B



Hund's rule tells us that orbitals closer to the nucleus are filled. Hund's rule tells us that orbitals closer to the nucleus are filled. Hund's rule tells us that orbitals closer to the nucleus are filled. Hund's rule tells us that orbitals closer to the nucleus are filled.

The following are the electron orbital diagrams for the second element, nitrogen and oxygen. Write the element's 2p electrons and explain to show Hund's rule. The 2p electrons are filled before the 2p electrons are filled.



Critical Thinking Questions

- How many "upspin" electrons are in a nitrogen atom? **1**
- Why does carbon's orbitals distribution is given as 1s² 2s² 2p²? Why not 1s² 2s² 2p²? **The 2p orbitals are already full. The 2s orbitals are already full. The 2p orbitals are already full.**
- Why do electrons fill orbitals for phosphorus? **1s² 2s² 2p² 3s² 3p² 4s² 4p² 5s² 5p² 6s² 6p² 7s² 7p² 8s² 8p² 9s² 9p² 10s² 10p² 11s² 11p² 12s² 12p² 13s² 13p² 14s² 14p² 15s² 15p² 16s² 16p² 17s² 17p² 18s² 18p² 19s² 19p² 20s² 20p² 21s² 21p² 22s² 22p² 23s² 23p² 24s² 24p² 25s² 25p² 26s² 26p² 27s² 27p² 28s² 28p² 29s² 29p² 30s² 30p² 31s² 31p² 32s² 32p² 33s² 33p² 34s² 34p² 35s² 35p² 36s² 36p² 37s² 37p² 38s² 38p² 39s² 39p² 40s² 40p² 41s² 41p² 42s² 42p² 43s² 43p² 44s² 44p² 45s² 45p² 46s² 46p² 47s² 47p² 48s² 48p² 49s² 49p² 50s² 50p² 51s² 51p² 52s² 52p² 53s² 53p² 54s² 54p² 55s² 55p² 56s² 56p² 57s² 57p² 58s² 58p² 59s² 59p² 60s² 60p² 61s² 61p² 62s² 62p² 63s² 63p² 64s² 64p² 65s² 65p² 66s² 66p² 67s² 67p² 68s² 68p² 69s² 69p² 70s² 70p² 71s² 71p² 72s² 72p² 73s² 73p² 74s² 74p² 75s² 75p² 76s² 76p² 77s² 77p² 78s² 78p² 79s² 79p² 80s² 80p² 81s² 81p² 82s² 82p² 83s² 83p² 84s² 84p² 85s² 85p² 86s² 86p² 87s² 87p² 88s² 88p² 89s² 89p² 90s² 90p² 91s² 91p² 92s² 92p² 93s² 93p² 94s² 94p² 95s² 95p² 96s² 96p² 97s² 97p² 98s² 98p² 99s² 99p² 100s² 100p² 101s² 101p² 102s² 102p² 103s² 103p² 104s² 104p² 105s² 105p² 106s² 106p² 107s² 107p² 108s² 108p² 109s² 109p² 110s² 110p² 111s² 111p² 112s² 112p² 113s² 113p² 114s² 114p² 115s² 115p² 116s² 116p² 117s² 117p² 118s² 118p² 119s² 119p² 120s² 120p² 121s² 121p² 122s² 122p² 123s² 123p² 124s² 124p² 125s² 125p² 126s² 126p² 127s² 127p² 128s² 128p² 129s² 129p² 130s² 130p² 131s² 131p² 132s² 132p² 133s² 133p² 134s² 134p² 135s² 135p² 136s² 136p² 137s² 137p² 138s² 138p² 139s² 139p² 140s² 140p² 141s² 141p² 142s² 142p² 143s² 143p² 144s² 144p² 145s² 145p² 146s² 146p² 147s² 147p² 148s² 148p² 149s² 149p² 150s² 150p² 151s² 151p² 152s² 152p² 153s² 153p² 154s² 154p² 155s² 155p² 156s² 156p² 157s² 157p² 158s² 158p² 159s² 159p² 160s² 160p² 161s² 161p² 162s² 162p² 163s² 163p² 164s² 164p² 165s² 165p² 166s² 166p² 167s² 167p² 168s² 168p² 169s² 169p² 170s² 170p² 171s² 171p² 172s² 172p² 173s² 173p² 174s² 174p² 175s² 175p² 176s² 176p² 177s² 177p² 178s² 178p² 179s² 179p² 180s² 180p² 181s² 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