

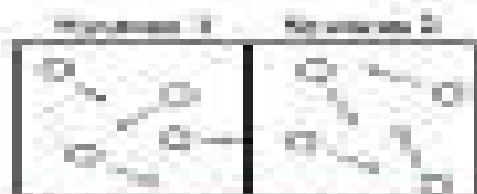
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v speed is $\frac{1}{2}mv^2$, the first being the mass and momentum being the change, directly connected to all of temperature increase. The Boltzmann distribution curve shows that the average velocity is small for the distribution's maximum value to change the kinetic energy increases.

3. Sketch a curve used to represent the relative velocities of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. The kinetic energy and mass of the particles are the same, but the Boltzmann distribution curve for System 1 is narrower and taller than the curve for System 2. This indicates that the particles in System 1 have a lower average kinetic energy and thus a lower temperature.
- b. Are the particles in System 1 moving faster or slower than the particles in System 2? Justify your answer in terms of molecular collisions.
 The particles in System 1 are moving faster than the particles in System 2. The particles in System 1 are at a higher temperature, which means they have a higher average kinetic energy. This is reflected in the Boltzmann distribution curve for System 1, which is shifted to the right of the curve for System 2, indicating that the particles in System 1 have a higher average speed.

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