

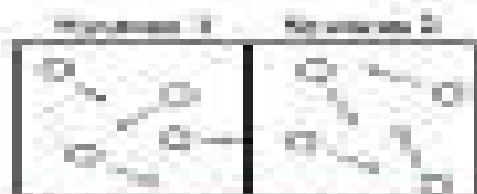
**Unit 6: Thermochemistry****6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



**Temperature:** is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with  $v$  is  $\frac{1}{2}mv^2$ , so the more of particles that move fast enough, the more kinetic energy we will see for a system. Therefore, the temperature of a system is proportional to the average kinetic energy of the particles. The Boltzmann-probability distribution curves show that the average kinetic energy of the particles is proportional to the temperature of the system.

3. Sketch a curve to represent the relative distribution of the particles in the following two gases, samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.  
System 1 is at a higher temperature than System 2. Temperature is a measure of the average kinetic energy of the particles in a system. The kinetic energy of a particle is  $\frac{1}{2}mv^2$ , so the more of particles that move fast enough, the more kinetic energy we will see for a system. Therefore, the temperature of a system is proportional to the average kinetic energy of the particles. The Boltzmann-probability distribution curves show that the average kinetic energy of the particles is proportional to the temperature of the system. In System 1, the particles are moving faster than in System 2, so the temperature of System 1 is higher than the temperature of System 2.
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.  
Thermal energy will be transferred from System 1 to System 2. Thermal energy is the total kinetic energy of the particles in a system. The kinetic energy of a particle is  $\frac{1}{2}mv^2$ , so the more of particles that move fast enough, the more kinetic energy we will see for a system. Therefore, the temperature of a system is proportional to the average kinetic energy of the particles. The Boltzmann-probability distribution curves show that the average kinetic energy of the particles is proportional to the temperature of the system. In System 1, the particles are moving faster than in System 2, so the temperature of System 1 is higher than the temperature of System 2. Thermal energy will be transferred from System 1 to System 2.

# Answer Key Unit 6 Thermochemistry

**Lei Shi**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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