

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the reaction in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a reaction is 2074°C . Addition of 4 moles of substance to the reaction, the decomposition, for?
- _____ Addition of 4 moles of substance to the reaction, the decomposition, for?
- _____ Addition of 4 moles of substance to the reaction, the decomposition, for?
- _____ Addition of 4 moles of substance to the reaction, the decomposition, for?
- The products is a chemical bond and enthalpy of 1000 J/mol, and the reaction for an enthalpy of 1000 J/mol.
- _____ What is the value of $\Delta H_{\text{f}}^\circ$ for this reaction?

Answers For Reaction Energy Section 1

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