

# Chemistry, Part 1

## Module 14 Notes



### Big Question

How are metals used in chemical reactions? and produced by chemical reactions?

### Big Idea

Using chemical analysis, you can identify the metals of a substance or pure metal and other substances. A metal is often used in manufacturing. It is used in construction or industrial analysis with chemical reactions.

### Learning Objectives

After completing this module, students will be able to:

- Use the law of conservation of mass to balance chemical reactions.
- Use chemical reactions to describe processes of materials.
- Compare different metals and reactions.
- Describe properties of different metals.

### Module 14 Objectives

**Learning Objectives** These are the main objectives of chemical reactions.

Fill in the blanks.

1. The law of conservation of mass states that the mass of the reactants is equal to the mass of the products.
2. The law of conservation of mass states that the mass of the reactants is equal to the mass of the products.
3. The law of conservation of mass states that the mass of the reactants is equal to the mass of the products.

### Chemical Reactions and Equations

For the first reaction, the reactants are iron and oxygen.

Chemical reactions are the chemical reactions that produce products.

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### Module 14

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# Answers For Chemistry Module 14

**Paul Collinson, Paul Collison**



## Answers For Chemistry Module 14:

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