

NATURAL SCIENCES

Gr. 9 - LESSON PLAN - MATTER & MATERIALS

TOPIC: COMPOUNDS		
Sub-topic: Names of Compounds		Duration: 1 Hour
CONTENT & CONCEPTS (CAPS p63)		
KEY CONCEPTS 1. Chemical reaction 2. Chemical formula 3. Names of compounds according to their elements.		
Specific Aims:	Specific Aim 1: 'Doing Science'	X
	Specific Aim 2: 'Knowing the subject content and making connections'	X
	Specific Aim 3: 'Understanding the uses of Science'	
LESSON OBJECTIVES: 1. Learners will be able to write the names of compounds as indicated on page 63 of the CAPS Document. 2. Learners will be able to make models of different compounds		
RESOURCES REQUIRED: • Periodic Table Poster • Siyavula Sasol Inzalo (Learner book) page 163		
TEACHING & LEARNING ACTIVITY: Introduction <u>Briefly review and revise concepts dealt with in Grade 8, focusing on compounds</u> As each element has a unique name for example: - Hydrogen - Lithium Remember the first 20 elements learnt in grade 8. How did you memorize them? Do you remember Mnemonics... He - H; He - He; Lies - Li; Because - Be; Boys - B; Can - C; Not - N; Operate - O; Fireplaces - F; New - Ne; Nation - Na; Might - Mg; Also - Al; Sign - Si; Peace - P; Security - S; Clause - Cl A - Ar; King - K; Can - Ca LEARNERS MAY RECITE THIS MNEMONIC AS A WAY OF MEMORIZING THE FIRST 20 ELEMENTS. <u>The teacher provides the following explanation to the learners</u> Each compound has a unique name. How is a compound formed? - A compound is formed when TWO or more different elements combine chemically. - Each atom in a compound are held by chemical bonds - Compounds are formed during chemical reaction (when elements/ compounds combine to form new substances) - Each compound can be written as a formula, for example, H₂O; NaCl; CuSO₄ What is a formula ? - This is a notation that shows which elements the compound is made up of. - It indicates the total number of atoms in one unit of the compound - It indicates the ratio in which atoms of different elements are bonded to each other - It gives the total number of atoms in one unit of a compound. For example, the formula for Copper Sulphate is CuSO ₄		

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Table of Contents 14 Grade 9 Ems

1. Understanding the eBook 14 Grade 9 Ems
 - The Rise of Digital Reading 14 Grade 9 Ems
 - Advantages of eBooks Over Traditional Books
2. Identifying 14 Grade 9 Ems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 14 Grade 9 Ems
 - User-Friendly Interface
4. Exploring eBook Recommendations from 14 Grade 9 Ems
 - Personalized Recommendations
 - 14 Grade 9 Ems User Reviews and Ratings
 - 14 Grade 9 Ems and Bestseller Lists
5. Accessing 14 Grade 9 Ems Free and Paid eBooks
 - 14 Grade 9 Ems Public Domain eBooks
 - 14 Grade 9 Ems eBook Subscription Services
 - 14 Grade 9 Ems Budget-Friendly Options
6. Navigating 14 Grade 9 Ems eBook Formats
 - ePub, PDF, MOBI, and More
 - 14 Grade 9 Ems Compatibility with Devices
 - 14 Grade 9 Ems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 14 Grade 9 Ems
 - Highlighting and Note-Taking 14 Grade 9 Ems
 - Interactive Elements 14 Grade 9 Ems
8. Staying Engaged with 14 Grade 9 Ems

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 14 Grade 9 Ems
- 9. Balancing eBooks and Physical Books 14 Grade 9 Ems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 14 Grade 9 Ems
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 14 Grade 9 Ems
 - Setting Reading Goals 14 Grade 9 Ems
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 14 Grade 9 Ems
 - Fact-Checking eBook Content of 14 Grade 9 Ems
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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