
OCR B Salters Chemistry

History of the atom correct answers Solid sphere

Johnson - Plum pudding (-ve electrons)

Rutherford - fired alpha particles at thin gold sheet, a few were deflected backwards - tiny +ve nucleus, surrounded by cloud of -ve electrons

Electrons in shells/ energy levels

Percentage yield equation correct answers $\text{actual yield} / \text{theoretical yield} \times 100$

Making new concentrations equation correct answers $\text{vol to use} = \text{final conc} / \text{initial conc} \times \text{vol required}$

Acid-alkali titrations correct answers Use phenolphthalein (pink $\rightarrow$ colourless)

Use methyl orange (yellow $\rightarrow$ red)

Things to remember in an ionic bonding diagram correct answers Added/lost electron as a dot/cross

Square brackets and overall charge

Properties of ionic compounds correct answers Conduct electricity when molten/(aq) - e^- are delocalised.

High melting points - strong ionic bonds require lots of energy to break

Soluble - ions are charged so are pulled apart by the polarity of water

Properties of simple covalent compounds correct answers Low m.p. - only need to overcome weak IMF

Don't conduct - no charge carriers

Usually insoluble - polar molecules are not attracted to the molecule

Properties of giant covalent molecules correct answers High m.p. - need to break strong bonds

Good conductors

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Max Parsonage



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