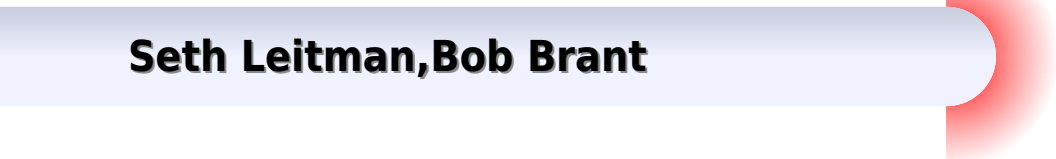




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web use the information in this chart to answer the following questions 1 if i have 125 grams of water at a temperature of 25.0 °C and heat it to 145.0 °C what will the change in enthalpy be 2 sketch the heating curve temperature vs time for the process in 1 3 i have 18.7 grams of an unknown liquid chemical at a temperature of 35.0 °C

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web sample problem calorimetry and enthalpy changes in an experiment 25.0 ml of 1.00 M HCl at 25.0 °C is added to 25.0 ml of 1.00 M NaOH at 25.0 °C in a foam cup calorimeter a reaction occurs and the temperature rises to 32.0 °C calculate the

enthalpy change in kJ for this reaction assume the densities of the solutions are 1.00 g/mL and that

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web before we practice calorimetry problems involving chemical reactions consider a simple example that illustrates the core idea behind calorimetry suppose we initially have a high temperature substance such as a hot piece of metal m and a low temperature substance such as cool water w answer the initial temperature of the copper

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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web in this set of practice questions we will go over the main types of questions on calorimetry including the heat capacity the heat of reaction finding the final temperature of a mixture constant pressure calorimetry and constant volume calorimetry a few important concepts and formulas you will need to solve these questions

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web calorimetry worksheet 1 if 0.315 moles of hexane C_6H_{14} is combusted in a bomb calorimeter containing 5.65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 55.4°C the specific heat capacity of water is 4.184 J/g°C h_{ms} t_{h} 5.650 grams H_2O 4.184 J/g°C 55.4°C h 1310 kJ

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web 6 a 218 g sample of steam at 121°C is cooled to ice at 14°C find the change in heat content of the system 7 in going from

ice at 34oc to steam at 138oc a sample of water absorbs $1.41 \times 10^5 \text{ J}$ find the mass of the sample answers $1.268 \times 10^4 \text{ J}$ $2.623 \times 10^4 \text{ J}$ $3.171 \times 10^4 \text{ J}$ $4.111 \times 10^5 \text{ J}$ $5.211 \times 10^5 \text{ J}$ 6

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