

SCHAU Unit 4 - Quiz – Equilibrium

Player(s): _____

Period(s): _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Relevant Equations:

$$K_c = \frac{[\text{products}]}{[\text{reactants}]} \quad ; \quad K_p = \frac{(P)^{\text{products}}}{(P)^{\text{reactants}}}$$

$$K_c = \frac{P_{\text{H}_2\text{O}}^2 P_{\text{CO}_2}}{P_{\text{CO}}^2 P_{\text{H}_2}^2} \quad ; \quad K_p = \frac{(P_{\text{H}_2\text{O}})^2 (P_{\text{CO}_2})}{(P_{\text{CO}})^2 (P_{\text{H}_2})^2}$$

gib (1 point) : 1.00 : gib (1) : gib (2) : gib (3) : gib (4) : gib (5) : gib (6)

1.00 : 1.00 : 1.00

Question 1:

1. Hydrogen sulfide, H_2S , is a slightly poisonous gas. At 1000 K, an equilibrium mixture contains 0.020 moles of hydrogen, H_2 , 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.020 atm. The reaction is:



If the volume of an 8.00 L container at 1000 K where a concentration of H_2 is present at equilibrium:

- What is the concentration of H_2 at equilibrium?
 Answer: H_2 is 0.0025 moles/L
 Answer: H_2 is 0.0025 moles/L
 H_2 is 0.0025 moles/L
- Calculate the value of K_c at equilibrium.
- Answer:
 What other species have chemical formula and a value of 0.0025 moles/L?
- Calculate:

	H_2	S_2	H_2S
Initial (moles/L)			
Change (moles/L)			
Final (moles/L)	0.0025	0.00125	0.0175

Final Equilibrium concentrations:

Answers For Sch4u Unit 4

SA Adler



Answers For Sch4u Unit 4:

Open court grade 4 unit 4 ,

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