

Experiment 10: Electrodes with a Standard Solution

From: *Basic Laboratory Experiments*, Experiment 10, 45-50

The purpose of this experiment is to determine the standard potential for several cells consisting of $\text{Zn} / \text{Zn}^{2+}$ and $\text{Cu} / \text{Cu}^{2+}$ electrodes. Electrode potentials are a tool to support calculations such as those involving Gibbs free energy, Gibbs free energy of activation, and other thermodynamic quantities. The standard potential for a half-cell is determined by measuring the cell potential of a cell consisting of the half-cell and a standard hydrogen electrode. The standard potential for a half-cell is determined by measuring the cell potential of a cell consisting of the half-cell and a standard hydrogen electrode. The standard potential for a half-cell is determined by measuring the cell potential of a cell consisting of the half-cell and a standard hydrogen electrode.

In this experiment, you will determine the standard potential for several electrodes and calculate the standard potential for several electrodes.

Objectives:

- to determine the standard potential for several electrodes
- to calculate the standard potential for several electrodes
- to determine the standard potential for several electrodes
- to determine the standard potential for several electrodes
- to calculate the standard potential for several electrodes

Materials:

- Standard solution
- Standard solution (100 mL)
- Standard solution
- Standard solution
- Standard solution
- Standard solution
- Standard solution
- Standard solution
- Standard solution

Procedure:

- Standard solution
- Standard solution
- Standard solution
- Standard solution

Results:

- Standard solution
- Standard solution

Discussion:

- This is a good starting point.
- This is approximately 100 mL of standard solution and the standard solution. The standard solution is approximately 100 mL of standard solution.

Answers For Lab 6b Heath Chemistry

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S. Coelho, Bin Tan, Robert Franzén**

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