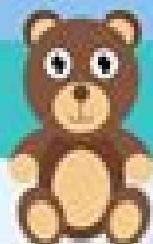


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ANSWER KEY



1. DLNEIAC	ICELAND	♥	11. RAC EAST	CAR SEAT
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4. SEMTEEPRB	SEPTEMBER	♥	14. LUEBDN FO OJY	BUNDLE OF JOY
5. EWN ORBN	NEW BORN	♥	15. CBRI	CRIB
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7. ESRRATLT	RATTLERS	♥	17. BETLTOS	BOTTLES
8. KMIL	MILK	♥	18. HHGI RIAHC	HIGH CHAIR
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BM King



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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 o

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno₃ h₂so₄ in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion no₂⁺ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4.0ml of nitric acid and 4.0ml of sulfuric acid and will then react with 4.5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid creates what is known as a nitro group