

**Question 1:** Water at  $20^{\circ}\text{C}$  falls from a height of 854 m. If the whole energy is used in increasing the temperature, find out the final temperature. Specific heat of water is  $4200 \text{ J/K kg}$ .

Given that

$$T_1 = 20^{\circ}\text{C} = 20 + 273 = 293 \text{ K}$$

$$\text{height} = h = 854 \text{ m}$$

$$\text{Specific heat of water} = 4200 \text{ J K}^{-1} \text{ kg}^{-1}$$

$$\text{Final temperature} = T_2 = ?$$

$$\text{P.E} = mgh \quad \text{--- (1)}$$

**Solution:**

We know that

$$\Delta Q = cm \Delta T \quad \text{--- (2)}$$

Whole energy is used in increasing the temperature (Given)

$$\Delta Q = \text{P.E}$$

From (1) and (2)

$$cm \Delta T = mgh$$

$$c \Delta T = gh$$

$$c (T_2 - T_1) = gh$$

$$T_2 - T_1 = \frac{gh}{c}$$

$$T_2 = \frac{gh}{c} + T_1$$

$$= \frac{9.8 \times 854}{4200} + 293$$

$$= \frac{8369.2}{4200} + 293 = 1.99 + 293$$

$$T_2 = 294.99 \text{ K} = 295 - 273^{\circ}\text{C}$$

$$T_2 = 22^{\circ}\text{C}$$

# Answer The Following Questions About The Thermodynamics Of

**Ying-Ying Zheng**



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## **Table of Contents Answer The Following Questions About The Thermodynamics Of**

1. Understanding the eBook Answer The Following Questions About The Thermodynamics Of
  - The Rise of Digital Reading Answer The Following Questions About The Thermodynamics Of
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer The Following Questions About The Thermodynamics Of
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer The Following Questions About The Thermodynamics Of
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer The Following Questions About The Thermodynamics Of
  - Personalized Recommendations
  - Answer The Following Questions About The Thermodynamics Of User Reviews and Ratings
  - Answer The Following Questions About The Thermodynamics Of and Bestseller Lists
5. Accessing Answer The Following Questions About The Thermodynamics Of Free and Paid eBooks
  - Answer The Following Questions About The Thermodynamics Of Public Domain eBooks
  - Answer The Following Questions About The Thermodynamics Of eBook Subscription Services
  - Answer The Following Questions About The Thermodynamics Of Budget-Friendly Options

6. Navigating Answer The Following Questions About The Thermodynamics Of eBook Formats
  - ePub, PDF, MOBI, and More
  - Answer The Following Questions About The Thermodynamics Of Compatibility with Devices
  - Answer The Following Questions About The Thermodynamics Of Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answer The Following Questions About The Thermodynamics Of
  - Highlighting and Note-Taking Answer The Following Questions About The Thermodynamics Of
  - Interactive Elements Answer The Following Questions About The Thermodynamics Of
8. Staying Engaged with Answer The Following Questions About The Thermodynamics Of
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer The Following Questions About The Thermodynamics Of
9. Balancing eBooks and Physical Books Answer The Following Questions About The Thermodynamics Of
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer The Following Questions About The Thermodynamics Of
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answer The Following Questions About The Thermodynamics Of
  - Setting Reading Goals Answer The Following Questions About The Thermodynamics Of
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer The Following Questions About The Thermodynamics Of
  - Fact-Checking eBook Content of Answer The Following Questions About The Thermodynamics Of
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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