

COVALENT BONDING

Draw Lewis structures for each molecule below. Circle all compounds that obey the octet rule and put a star next to all compounds that violate the octet rule. Indicate the bonding regions and unshared pairs for the central atom (or least electronegative excluding hydrogen) and tell how many sigma and pi bonds each compound has.

Molecule	Structure	Bonding Regions	Unshared Pairs	Sigma Bonds	Pi Bonds
SCl ₄		6	0	6	0
NO	N≡N-O:	2	0	2	2
NO ₂	$\cdot\ddot{O}-\ddot{N}=\ddot{O}\cdot \leftrightarrow \cdot\ddot{O}=\ddot{N}-\ddot{O}\cdot$	2	1/2	2	1
[NO ₂] ⁺	$[\cdot\ddot{O}-\ddot{N}=\ddot{O}]^+ \leftrightarrow [\cdot\ddot{O}=\ddot{N}-\ddot{O}]^+$	2	1	2	1
BBr ₃		3	0	3	0
PH ₃		3	1	3	0
[AsF ₃] ⁺		4	1	4	0
[SO ₃] ²⁻		4	0	4	0

molecular geometry

Octahedral

linear

bent

trigonal planar

trigonal pyramidal

See saw

tetrahedral

hybridization

sp³d²

sp

sp²

sp²

sp³

sp³d

sp³

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