VOLUMETRIC ANALYSIS

A IM

Determine the concentration of aqueous hydrogen peroxide and the percentage purity of oxalate in an unknown oxalate sample by titration with potassium permanganate.

INTRODUCTION

Volumetric analysis is a quantitative analysis process whereby volumes are measured. A volume of reagent known as the standard solution of known concentration is chemically reacted with a solution of unknown concentration in order to determine the concentration of the unknown. The solution that is used as standard is called the titrant which is delivered from a burette and is added until the reaction is completed. The end point of the titration is noted by observing the change of colour of an indicator. In a direct titration, the titrant is added to analyte until the end point is observed whereas in a back titration, a known excess of a standard reagent is added to the analyte. A second standard reagent is then used to titrate the excess of the first reagent.

All the calculations in determining the molarity and average concentration for all equations involve the formula:

$$M_BV_Bb$$

The equation for the standardization experiment is given in the ionic form as:

 $5C_2O_0^2 + 2MnO^6 + 16H^2 - 10CO_2 + 2Mn^2 + 8H_2O$

The back titration is explained by the equations:

5NO2+2MnO+6H - 5NO3+2Mn2+3H2O

APPARATUS AND REAGENTS

1.6g of Potassium permanganate

Boiling distilled water

Distilled water

Erlenmeyer flask

Rubber stopper

Sodium oxalate

Sulphuric acid

Source of heat

Burette

Hydrogen peroxide

A Volumetric Analysis Lab Report Answers

John E. Havel

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