
AAMC MCAT Practice Exam 1 Questions and Answers

In the chromatography of the reaction mixture, water absorbed on cellulose functioned as the stationary phase. What was the principal factor determining the migration of individual components in the sample?

- A) Hydrogen bonding
- B) Solute concentration
- C) Stationary phase concentration
- D) Thickness of paper - Answer - A) Hydrogen bonding

The answer to this question is A because the relative amount of hydrogen bonding to the stationary phase will determine the relative rate of migration of the various components in the sample.

What assumption is being made if scientists conclude that aspartic acid was formed by the prebiological synthesis in the passage?

- A) Aspartic acid is unstable at temperatures below 150 degrees C.
- B) All of the malic acid underwent the dehydration reaction to form fumaric/maleic acid.
- C) Compound A and cyanide were available on primitive Earth.
- D) The reaction between ammonia and fumaric acid was catalyzed by the presence of water. - Answer - C) Compound A and cyanide were available on primitive Earth.

The answer to this question is C since, in order for the experimental reaction sequence to be relevant to the primordial formation of aspartic acid, the starting materials used (Compound A and cyanide) are assumed to have been available.

According to the developed chromatography plate shown below, what is the approximate R_f value of aspartic acid?

aspartic acid = 2
solvent front = 10

- A) 0.20
- B) 0.50
- C) 5
- D) 10 - Answer - The answer to this question is A (0.20) because R_f is the ratio of the distance travelled by the analyte relative to the solvent front during a chromatographic separation. Aspartic acid travelled two units, while the solvent front travelled ten units, giving an R_f of $2/10 = 0.20$ for aspartic acid.

Which of the following statements does NOT correctly describe the

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Table of Contents Aamc Test 10 Mcat

1. Understanding the eBook Aamc Test 10 Mcat
 - The Rise of Digital Reading Aamc Test 10 Mcat
 - Advantages of eBooks Over Traditional Books
2. Identifying Aamc Test 10 Mcat
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Aamc Test 10 Mcat
 - User-Friendly Interface
4. Exploring eBook Recommendations from Aamc Test 10 Mcat
 - Personalized Recommendations
 - Aamc Test 10 Mcat User Reviews and Ratings
 - Aamc Test 10 Mcat and Bestseller Lists
5. Accessing Aamc Test 10 Mcat Free and Paid eBooks
 - Aamc Test 10 Mcat Public Domain eBooks

- Aamc Test 10 Mcat eBook Subscription Services
- Aamc Test 10 Mcat Budget-Friendly Options
- 6. Navigating Aamc Test 10 Mcat eBook Formats
 - ePub, PDF, MOBI, and More
 - Aamc Test 10 Mcat Compatibility with Devices
 - Aamc Test 10 Mcat Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Aamc Test 10 Mcat
 - Highlighting and Note-Taking Aamc Test 10 Mcat
 - Interactive Elements Aamc Test 10 Mcat
- 8. Staying Engaged with Aamc Test 10 Mcat
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Aamc Test 10 Mcat
- 9. Balancing eBooks and Physical Books Aamc Test 10 Mcat
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Aamc Test 10 Mcat
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Aamc Test 10 Mcat
 - Setting Reading Goals Aamc Test 10 Mcat
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Aamc Test 10 Mcat
 - Fact-Checking eBook Content of Aamc Test 10 Mcat
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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