

A Voyage through Equations ANSWER KEY

Section 1: Identify the type of reaction

- 1) $\text{Na}_3\text{PO}_4 + 3 \text{KOH} \rightarrow 3 \text{NaOH} + \text{K}_3\text{PO}_4$ DOUBLE DISPLACEMENT
- 2) $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{LiCl}$ DOUBLE DISPLACEMENT
- 3) $\text{C}_6\text{H}_{12} + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$ COMBUSTION
- 4) $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ SINGLE DISPLACEMENT
- 5) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ DECOMPOSITION
- 6) $\text{P}_4 + 3 \text{O}_2 \rightarrow 2 \text{P}_2\text{O}_3$ SYNTHESIS
- 7) $2 \text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be}(\text{NO}_3)_2 + 2 \text{RbF}$ DOUBLE DISPLACEMENT
- 8) $2 \text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu}(\text{NO}_3)_2 + 2 \text{Ag}$ SINGLE DISPLACEMENT
- 9) $\text{C}_2\text{H}_6\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$ COMBUSTION
- 10) $2 \text{C}_2\text{H}_6 + \text{Fe} \rightarrow \text{Fe}(\text{C}_2\text{H}_5)_2$ SYNTHESIS
- 11) $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + 3 \text{Cl}_2$ SINGLE DISPLACEMENT
- 12) $2 \text{Hg}_2 + 8 \text{HClO}_2 \rightarrow 2 \text{HgCl}_2 + 8 \text{HCl}$ DOUBLE DISPLACEMENT
- 13) $\text{O}_3 \rightarrow \text{O} + \text{O}_2$ DECOMPOSITION
- 14) $2 \text{HCl} + 2 \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ DECOMPOSITION

Section 2: Practicing equation balancing

- 1) $2 \text{CuCl} + 2 \text{Fe} + 2 \text{HCl} + 2 \text{HCl}$
- 2) $2 \text{Na} + 2 \text{HClO}_2 + 2 \text{HCl} + 2 \text{HClO}_2$
- 3) $2 \text{HCl} + 2 \text{HCl} + 2 \text{HCl} + 2 \text{HCl}$
- 4) $2 \text{HClO}_2 + 2 \text{HCl} + 2 \text{HCl}$
- 5) $2 \text{HCl} + 2 \text{HClO}_2 + 2 \text{HCl} + 2 \text{HClO}_2$
- 6) $2 \text{HCl} + 2 \text{HCl} + 2 \text{HCl} + 2 \text{HClO}_2$
- 7) $2 \text{HCl} + 2 \text{HCl} + 2 \text{HCl}$
- 8) $2 \text{HCl} + 2 \text{HCl} + 2 \text{HCl} + 2 \text{HCl}$
- 9) $2 \text{HCl} + 2 \text{HCl} + 2 \text{HClO}_2$
- 10) $2 \text{HCl} + 2 \text{HCl} + 2 \text{HCl}$

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Charles Knight



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