

Researched expected results

Test 1

	Magnesium chloride	Calcium bromide	Strontium chloride	Barium chloride
Initial	Colourless solution	Colourless solution	Colourless solution	Colourless solution
After adding 10 drops of 0.5M NaOH solution	Slightly white precipitate	Slightly white precipitate	Slightly white precipitate	Colourless solution
After adding excess NaOH	White precipitate	Slightly white precipitate	Slightly white precipitate	Colourless solution
Reaction and Product formed	$\text{MgCl}_2 + 2\text{NaOH} \rightarrow \text{Mg(OH)}_2(\text{s}) + 2\text{NaCl}$	$\text{CaBr}_2 + 2\text{NaOH} \rightarrow \text{Ca(OH)}_2(\text{s}) + 2\text{NaBr}$	$\text{SrCl}_2 + 2\text{NaOH} \rightarrow \text{Sr(OH)}_2(\text{s}) + 2\text{NaCl}$	$\text{BaCl}_2 + \text{NaOH} \rightarrow \text{Ba(OH)}_2 + \text{NaCl}$

Test 2

	Magnesium chloride	Calcium bromide	Strontium chloride	Barium chloride
Initial	Colourless solution	Colourless solution	Colourless solution	Colourless solution
After adding 10 drops of 1.0M H ₂ SO ₄ solution	Slight white precipitate	Slight white precipitate	white precipitate	white precipitate
After adding excess H ₂ SO ₄	Colourless solution	Slight white precipitate	white precipitate	white precipitate
Reaction and Product formed	$\text{MgCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + 2\text{HCl}$	$\text{CaBr}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{SO}_2 + \text{Br}_2 + \text{CaSO}_4(\text{s})$	$\text{SrCl}_2 + 2\text{H}_2\text{SO}_4 \rightarrow \text{Sr(HSO}_4)_2(\text{s}) + 2\text{HCl}$	$\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4(\text{s}) + \text{HCl}$

Test 3

Observation of damp red litmus paper being held over the test tube containing a heated mixture of 10 drops of 0.1M ammonium chloride solution and 10 drops 0.5M sodium hydroxide solution.	Litmus paper turns blue
Reaction and product formed	$\text{NH}_4\text{Cl} + \text{NaOH} \rightarrow \text{NH}_3 + \text{NaCl} + \text{H}_2\text{O}$
Reason	Ammonium ions can be recognised in a solution. If ammonium ions are present when added to sodium hydroxide, they will be converted to ammonia gas and/or turn red litmus paper blue.

Test 4

Observation of red litmus paper with the fumes of 1cm ³ solution of 0.5M sodium hydroxide	Turns damp red litmus paper blue
Observations made after a few minutes of dampening red litmus paper with deionised water and placing it beside filter paper soaked in 1M ammonia solution inside a closed petri dish	Turns damp red litmus paper blue

Test 5

Solution	Volume (cm ³)	Molarity (M)	Observation
Limewater	1	0.02	Colourless solution
$\text{HCl} + \text{Na}_2\text{CO}_3 \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$	3	1.00	After adding hydrochloric acid and sodium carbonate to the limewater, the solution went from colourless to cloudy.

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