

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete chemical equation.

Vocabulary

- 1. chemical equation
- 2. chemical equation
- 3. reactant
- 4. coefficient
- 5. balanced equation

Part A Completion

Write the coefficient for each in these chemical equations, and balance and label every chemical reaction. Complete the equation below.

A chemical equation is always represented by a balanced _____. The chemical formula for a chemical compound is _____. The law of conservation of mass requires that the _____. An equation is written the way described in _____. A chemical equation is written as a _____. The steps for writing a balanced equation are given _____. In a chemical equation, the coefficients are the large number of atoms of an _____. The equation is balanced if the number of atoms of each element is the same on both sides of the equation. The equation is balanced if the number of atoms of each element is the same on both sides of the equation.

Chemical equations are written in a way that shows the reaction as a chemical _____. The equation is balanced if the number of atoms of each element is the same on both sides of the equation. The equation is balanced if the number of atoms of each element is the same on both sides of the equation. The equation is balanced if the number of atoms of each element is the same on both sides of the equation.

1. equation
2. chemical
3. conservation
4. balanced
5. coefficient
6. equation
7. balanced
8. _____
9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. _____
19. _____
20. _____

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Ralph Thomas Myers



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