

COMMON TYPES OF CHEMICAL REACTIONS

SYNTHESIS

Multiple reactants, which can be simple elements or compounds, combine together to form a single compound.



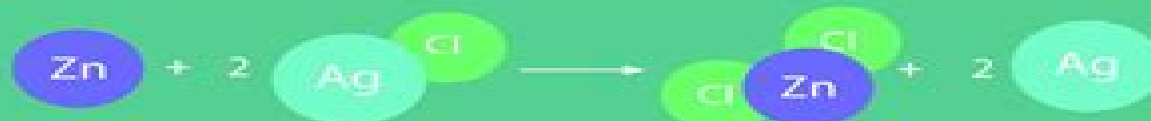
DECOMPOSITION

A compound breaks down into two or more simpler substances. Decomposition reactions are classified into thermal, electrolytic, and photo.



SINGLE-REPLACEMENT

One element is substituted for another element in a compound, generating a new compound and a pure element.



DOUBLE-REPLACEMENT

Two ionic compounds exchange cations or anions to form two new compounds. Forming a precipitate can help drive the reaction to the right.



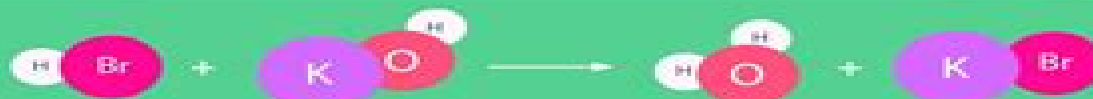
COMBUSTION

A fuel, typically a hydrocarbon, reacts with oxygen gas to form carbon dioxide and water, which generates heat and light.



NEUTRALIZATION

Acid and base combine to neutralize each other in aqueous solutions, usually resulting in a salt and water.



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