

12.3 Limiting Reagent and Percent Yield >

Limiting and Excess Reagents

Experimental Conditions		
	Reactants	Products
Before reaction	 2 molecules N_2  3 molecules H_2	0 molecules NH_3
After reaction	 1 molecule N_2  0 molecules H_2	 2 molecules NH_3

- In this reaction, only the hydrogen is completely used up.
- H_2 is the **limiting reagent**, or the reactant that determines the amount of product that can be formed by a reaction.

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Ray Dexter



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About the Author: Adrian Dingle is a chemistry educator and author with 24 years of experience teaching in the United States and the United Kingdom. He is the creator of the award-winning chemistry website www.adriandingleschemistrypages.com. The focus of Mr. Dingle's teaching career has been on preparing students for standardized tests: AP and SAT tests in the United States, GCSE's and A levels in the United Kingdom, and International Baccalaureate in both countries. An Englishman, he lives in Atlanta, Georgia, where he teaches at The Westminster Schools. He holds a B.Sc. Honors Chemistry and a Postgraduate Certificate in Education, both from the University of Exeter in England. In addition to writing this Crash Course, Mr. Dingle has written *The Periodic Table Elements With Style*, *How To Make A Universe With 92 Ingredients*, and *SAT Chemistry Crash Course*. He is the 2011 winner of the School Library Association of the UK's Information Book Award and in 2012 was honored with the prestigious literary prize *Wissenschaftsbuch des Jahres* sponsored by the Austrian Ministry of Science and Research.

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