
Reinforce and Review

Lesson Worksheets

Copy and distribute the Lesson 12.3 worksheets in CK-12 Chemistry –Intermediate Workbook. Ask students to complete the worksheets to reinforce lesson content.

Lesson Review Questions

Have students answer the Lesson Review Questions at the end of Lesson 12.3 in CK-12 Chemistry –Intermediate FlexBook® resource.

Sample Answers

Practice Problems

1.

$$1.30 \text{ mol } C_3H_8 \times \frac{5 \text{ mol } O_2}{1 \text{ mol } C_3H_8} = 6.50 \text{ mol } O_2 \text{ needed}$$

Since the moles of O_2 needed to fully react with all of the C_3H_8 is more than the amount of O_2 present (6.00 mol), O_2 is the limiting reactant.

2.

$$8.00 \text{ g } Zn \times \frac{1 \text{ mol } Zn}{65.38 \text{ g } Zn} \times \frac{2 \text{ mol } HCl}{1 \text{ mol } Zn} \times \frac{36.46 \text{ g } HCl}{1 \text{ mol } HCl} = 8.92 \text{ g } HCl \text{ needed}$$

Since there is less HCl present (8.00 g) than what is needed to react with all of the Zn, the HCl is the limiting reactant. This problem can also be solved by calculating the moles of Zn and the moles of HCl originally present. Compare the number of moles, taking note of the 1:2 ratio.

$$8.00 \text{ g } Zn \times \frac{1 \text{ mol } Zn}{65.38 \text{ g } Zn} = 0.122 \text{ mol } Zn$$
$$8.00 \text{ g } HCl \times \frac{1 \text{ mol } Zn}{36.46 \text{ g } HCl} = 0.219 \text{ mol } HCl$$

The moles of HCl is not double the moles of Zn, as required by the balanced equation. Therefore, HCl is the limiting reactant.

3.

- The balanced equation requires twice as many moles of Na as Cl_2 . Sodium (Na) is the limiting reactant.
- The 3.40 moles of Na can be converted to moles of Cl_2 that would react.

$$3.40 \text{ mol } Na \times \frac{1 \text{ mol } Cl_2}{2 \text{ mol } Na} = 1.70 \text{ mol } Cl_2 \text{ reacts}$$

The excess Cl_2 is determined by subtraction. $4.20 \text{ mol} - 1.70 \text{ mol} = 2.50 \text{ mol}$ excess Cl_2

- $3.40 \text{ mol } Na \times \frac{2 \text{ mol } NaCl}{1 \text{ mol } Na} = 3.40 \text{ mol } NaCl$

4. First, the limiting reactant must be determined.

$$35.0 \text{ g } N_2 \times \frac{1 \text{ mol } N_2}{28.02 \text{ g } N_2} \times \frac{3 \text{ mol } H_2}{1 \text{ mol } N_2} \times \frac{2.02 \text{ g } H_2}{1 \text{ mol } H_2} = 7.57 \text{ g } H_2 \text{ needed}$$

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S. Rao Vallabhaneni



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