

9 Calculations from chemical equations

Answers to end-of-chapter questions

Pages 163–164 Questions

1. molar mass of $\text{Fe}_2(\text{SO}_4)_3 = [2 \times 55.8] + 3 \times [32.1 + (4 \times 16.0)] = 399.9 \text{ g mol}^{-1}$
molar mass of $\text{Fe}(\text{OH})_3 = 55.8 + (3 \times 17.0) = 106.8 \text{ g mol}^{-1}$
amount of $\text{Fe}_2(\text{SO}_4)_3$ (in moles) = mass/molar mass = $12.7 \text{ g} / 399.9 \text{ g mol}^{-1} = 0.03176 \text{ mol}$
moles $\text{Fe}(\text{OH})_3$ = moles $\text{Fe}_2(\text{SO}_4)_3 \times$ number of $\text{Fe}(\text{OH})_3$ in equation/number of $\text{Fe}_2(\text{SO}_4)_3$ in equation
= $0.03176 \text{ mol} \times 2 = 0.06352 \text{ mol}$
mass of $\text{Fe}(\text{OH})_3$ = moles $\times$ molar mass = $0.06352 \text{ mol} \times 106.8 \text{ g mol}^{-1} = 6.78 \text{ g}$

[e] It is a good idea to work out the molar masses of the two substances involved before starting the rest of the calculation.

2. molar mass of $\text{AgNO}_3 = 107.9 + 14.0 + (3 \times 16.0) = 169.9 \text{ g mol}^{-1}$
amount (in moles) of AgNO_3 = mass/molar mass = $12.6 \text{ g} / 169.9 \text{ g mol}^{-1} = 0.07416 \text{ mol}$
ratio of $\text{Cu}:\text{AgNO}_3 = 1:2$
so, amount of copper = $\frac{1}{2} \times 0.07416 = 0.03708 \text{ mol}$
mass of copper needed = moles $\times$ molar mass
= $0.03708 \text{ mol} \times 63.5 \text{ g mol}^{-1} = 2.35 \text{ g}$

[e] The answer 2.35 g was obtained keeping all the numbers on the calculator during the calculation. If the rounded up value of 0.0371 mol is used, the answer 2.36 g is obtained. Either would score full marks.

A common error is to calculate the molar mass of AgNO_3 as $2 \times 170 \text{ g mol}^{-1}$ because there are 2 moles of it in the equation. This is wrong. The number of moles of a substance depends only on its mass and not on the reaction. The stoichiometry comes into play when moles of one substance are converted to moles of another substance. This type of calculation can be done either by mass ratio or by converting to moles, then using the reaction stoichiometry and finally converting back to mass. The second method fits all types of calculation, so is the better one to use.

3. molar mass of $\text{NaNO}_3 = 23.0 + 14.0 + (3 \times 16.0) = 85.0 \text{ g mol}^{-1}$
amount of NaNO_3 = mass/molar mass = $33.3 \text{ g} / 85.0 \text{ g mol}^{-1} = 0.392 \text{ mol}$
ratio of $\text{O}_2:\text{NaNO}_3 = 1:2$
so, amount of oxygen = $\frac{1}{2} \times 0.392 = 0.196 \text{ mol}$
volume of oxygen = moles $\times$ molar volume
= $0.196 \text{ mol} \times 25.0 \text{ dm}^3 \text{ mol}^{-1}$
= 4.90 dm^3
4. ratio of $\text{NH}_3:\text{H}_2 = 2:3$

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Paul Yates



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Table of Contents 122 Chemical Calculations Answer Key

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 - The Rise of Digital Reading 122 Chemical Calculations Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 122 Chemical Calculations Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 122 Chemical Calculations Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 122 Chemical Calculations Answer Key
 - Personalized Recommendations
 - 122 Chemical Calculations Answer Key User Reviews and Ratings
 - 122 Chemical Calculations Answer Key and Bestseller Lists
5. Accessing 122 Chemical Calculations Answer Key Free and Paid eBooks

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 - 122 Chemical Calculations Answer Key Budget-Friendly Options
6. Navigating 122 Chemical Calculations Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - 122 Chemical Calculations Answer Key Compatibility with Devices
 - 122 Chemical Calculations Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 122 Chemical Calculations Answer Key
 - Highlighting and Note-Taking 122 Chemical Calculations Answer Key
 - Interactive Elements 122 Chemical Calculations Answer Key
8. Staying Engaged with 122 Chemical Calculations Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 122 Chemical Calculations Answer Key
9. Balancing eBooks and Physical Books 122 Chemical Calculations Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 122 Chemical Calculations Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 122 Chemical Calculations Answer Key
- Setting Reading Goals 122 Chemical Calculations Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 122 Chemical Calculations Answer Key
- Fact-Checking eBook Content of 122 Chemical Calculations Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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