

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, $\text{KClO}_3(\text{s})$. Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{23} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



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Peipei Pang



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Table of Contents 12 Stoichiometry Guided And Study Workbook Answers 235324

1. Understanding the eBook 12 Stoichiometry Guided And Study Workbook Answers 235324
 - The Rise of Digital Reading 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Advantages of eBooks Over Traditional Books
2. Identifying 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 12 Stoichiometry Guided And Study Workbook Answers 235324
 - User-Friendly Interface
4. Exploring eBook Recommendations from 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Personalized Recommendations
 - 12 Stoichiometry Guided And Study Workbook Answers 235324 User Reviews and Ratings

- 12 Stoichiometry Guided And Study Workbook Answers 235324 and Bestseller Lists
- 5. Accessing 12 Stoichiometry Guided And Study Workbook Answers 235324 Free and Paid eBooks
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- 6. Navigating 12 Stoichiometry Guided And Study Workbook Answers 235324 eBook Formats
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 - 12 Stoichiometry Guided And Study Workbook Answers 235324 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Highlighting and Note-Taking 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Interactive Elements 12 Stoichiometry Guided And Study Workbook Answers 235324
- 8. Staying Engaged with 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 12 Stoichiometry Guided And Study Workbook Answers 235324
- 9. Balancing eBooks and Physical Books 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 12 Stoichiometry Guided And Study Workbook Answers 235324
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Setting Reading Goals 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Fact-Checking eBook Content of 12 Stoichiometry Guided And Study Workbook Answers 235324
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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