

Consequences of Schottky defect

- As the number of ions decreases, mass decreases. However, volume remains unchanged. Hence, the density of a substance decreases.
- The number of missing cations and anions is equal, the electrical neutrality of the compound is preserved.

This defect is found in ionic crystals such as NaCl, AgBr and KCl.

iv. Frenkel defect : Frenkel defect arises when an ion of an ionic compound is missing from its regular lattice site and occupies interstitial position between lattice points as shown in Fig. 1.16.

The cations are usually smaller than anions. It is, therefore, more common to find the cations occupying interstitial sites. It is easier for the smaller cations to accommodate the interstitial spaces.

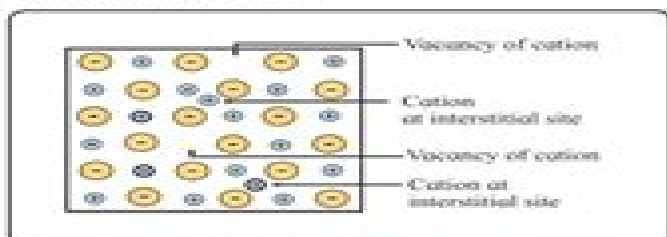


Fig. 1.16 : Frenkel defect

Do you know ?

Frenkel defect is not found in pure alkali metal halides because cations of alkali metals due to large size cannot occupy interstitial space.



It is important to note that the smaller cation is displaced from its normal site to an interstitial space. It, therefore, creates a vacancy defect at its original position and interstitial defect at its new location in the same crystal. Frenkel defect can be regarded as the combination of vacancy defect and interstitial defect.

Conditions for the formation of Frenkel defect

- Frenkel defect occurs in ionic compounds with large difference between sizes of cation and anion.
- The ions of ionic compounds must be having low coordination number.

Consequences of Frenkel defect

- As no ions are missing from the crystal lattice as a whole, the density of solid and its chemical properties remain unchanged.
- The crystal as a whole remains electrically neutral because the equal numbers of cations and anions are present.

This defect is found in ionic crystals like ZnS, AgCl, AgBr, AgI, CaF₂.

b. Impurity defect : Impurity defect arises when foreign atoms, that is, atoms different from the host atoms, are present in the crystal lattice. There are two kinds of impurity defects : Substitutional and interstitial impurity defects.

i. Substitutional impurity defect : In this defect, the foreign atoms are found at the lattice sites in place of host atoms. The regular atoms are displaced from their lattice sites by impurity atoms.

For example :

- **Solid solutions of metals (alloys) :** Brass is an alloy of Cu and Zn. In brass, host Cu atoms are replaced by impurity of Zn atoms. The Zn atoms occupy regular sites of Cu atoms as shown in Fig. 1.17.

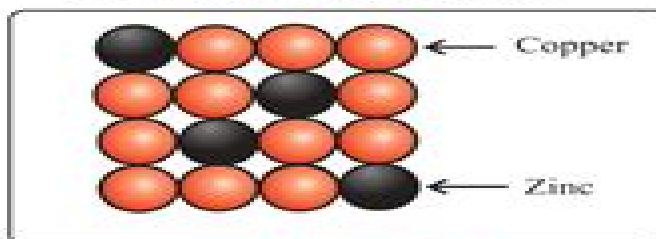


Fig. 1.17 : Brass

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Table of Contents 12th Std Chemistry Lecture Notes Maharashtra Board

1. Understanding the eBook 12th Std Chemistry Lecture Notes Maharashtra Board
 - The Rise of Digital Reading 12th Std Chemistry Lecture Notes Maharashtra Board
 - Advantages of eBooks Over Traditional Books
2. Identifying 12th Std Chemistry Lecture Notes Maharashtra Board
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 12th Std Chemistry Lecture Notes Maharashtra Board
 - User-Friendly Interface
4. Exploring eBook Recommendations from 12th Std Chemistry Lecture Notes Maharashtra Board
 - Personalized Recommendations

- 12th Std Chemistry Lecture Notes Maharashtra Board User Reviews and Ratings
- 12th Std Chemistry Lecture Notes Maharashtra Board and Bestseller Lists
- 5. Accessing 12th Std Chemistry Lecture Notes Maharashtra Board Free and Paid eBooks
 - 12th Std Chemistry Lecture Notes Maharashtra Board Public Domain eBooks
 - 12th Std Chemistry Lecture Notes Maharashtra Board eBook Subscription Services
 - 12th Std Chemistry Lecture Notes Maharashtra Board Budget-Friendly Options
- 6. Navigating 12th Std Chemistry Lecture Notes Maharashtra Board eBook Formats
 - ePub, PDF, MOBI, and More
 - 12th Std Chemistry Lecture Notes Maharashtra Board Compatibility with Devices
 - 12th Std Chemistry Lecture Notes Maharashtra Board Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 12th Std Chemistry Lecture Notes Maharashtra Board
 - Highlighting and Note-Taking 12th Std Chemistry Lecture Notes Maharashtra Board
 - Interactive Elements 12th Std Chemistry Lecture Notes Maharashtra Board
- 8. Staying Engaged with 12th Std Chemistry Lecture Notes Maharashtra Board
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 12th Std Chemistry Lecture Notes Maharashtra Board
- 9. Balancing eBooks and Physical Books 12th Std Chemistry Lecture Notes Maharashtra Board
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 12th Std Chemistry Lecture Notes Maharashtra Board
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 12th Std Chemistry Lecture Notes Maharashtra Board
 - Setting Reading Goals 12th Std Chemistry Lecture Notes Maharashtra Board
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 12th Std Chemistry Lecture Notes Maharashtra Board
 - Fact-Checking eBook Content of 12th Std Chemistry Lecture Notes Maharashtra Board

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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