

13.4 Practice Problems

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1. A gas cylinder contains a mixture of gases at a pressure of 1.00 atm and a temperature of 27.0°C. The gas is expanded to a pressure of 0.500 atm and a temperature of 12.0°C. Calculate the change in internal energy of the gas.
2. An air conditioner cools a room at 27.0°C to 19.0°C. Calculate the change in entropy of the air in the room. Assume the room is a closed system.
3. A thermodynamic cycle is performed on a gas. The cycle consists of four states: (1) 1.00 mol of gas at 1.00 atm and 27.0°C, (2) 1.00 mol of gas at 1.00 atm and 12.0°C, (3) 1.00 mol of gas at 0.500 atm and 12.0°C, and (4) 1.00 mol of gas at 0.500 atm and 27.0°C. Calculate the net work done by the gas during the cycle.
4. A gas cylinder contains a mixture of gases at a pressure of 1.00 atm and a temperature of 27.0°C. The gas is expanded to a pressure of 0.500 atm and a temperature of 12.0°C. Calculate the change in enthalpy of the gas.
5. A thermodynamic cycle is performed on a gas. The cycle consists of four states: (1) 1.00 mol of gas at 1.00 atm and 27.0°C, (2) 1.00 mol of gas at 1.00 atm and 12.0°C, (3) 1.00 mol of gas at 0.500 atm and 12.0°C, and (4) 1.00 mol of gas at 0.500 atm and 27.0°C. Calculate the net heat added to the gas during the cycle.
6. A gas cylinder contains a mixture of gases at a pressure of 1.00 atm and a temperature of 27.0°C. The gas is expanded to a pressure of 0.500 atm and a temperature of 12.0°C. Calculate the change in Gibbs free energy of the gas.
7. A thermodynamic cycle is performed on a gas. The cycle consists of four states: (1) 1.00 mol of gas at 1.00 atm and 27.0°C, (2) 1.00 mol of gas at 1.00 atm and 12.0°C, (3) 1.00 mol of gas at 0.500 atm and 12.0°C, and (4) 1.00 mol of gas at 0.500 atm and 27.0°C. Calculate the net change in Gibbs free energy of the gas during the cycle.

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Albert A Gayle



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Mathematics, 1989 *Cumulated Index Medicus*, 1994

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