

# Gravimetric Analysis of Chloride in Solution

Experiment 10  
Lab 10  
Date: \_\_\_\_\_

## Introduction:

This experiment involves the gravimetric analysis of a solution containing chloride ions. The principle of gravimetric analysis is based on the fact that the mass of a substance is proportional to the number of moles of the substance. In this experiment, the chloride ions in the solution are precipitated as silver chloride ( $\text{AgCl}$ ) by adding a known volume of a silver nitrate solution. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a dry environment to avoid the presence of water, which can interfere with the precipitation reaction. The experiment is performed in a dry environment to avoid the presence of water, which can interfere with the precipitation reaction. The experiment is performed in a dry environment to avoid the presence of water, which can interfere with the precipitation reaction.

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## Procedure:

The procedure for this experiment is as follows: 1. Weigh a known mass of the sample (e.g., 0.1000 g) and transfer it to a beaker. 2. Add a known volume of water (e.g., 100 mL) to the beaker. 3. Add a known volume of silver nitrate solution (e.g., 10 mL) to the beaker. 4. Stir the mixture thoroughly. 5. Filter the mixture through a filter paper. 6. Wash the filter paper with distilled water. 7. Dry the filter paper in a desiccator. 8. Weigh the dried precipitate. 9. Calculate the mass of chloride ions in the original sample.

## Data Analysis:

|                      | Sample 1 | Sample 2 |
|----------------------|----------|----------|
| Mass of sample       | 0.1000 g | 0.1000 g |
| Mass of precipitate  | 0.1500 g | 0.1500 g |
| Mass of chloride ion | 0.0500 g | 0.0500 g |

The mass of chloride ion in the sample is calculated as follows:  $m_{\text{Cl}^-} = m_{\text{AgCl}} \times \frac{m_{\text{Cl}^-}}{m_{\text{AgCl}}}$

For Sample 1:  $m_{\text{Cl}^-} = 0.1500 \text{ g} \times \frac{35.45}{143.32} = 0.0375 \text{ g}$

For Sample 2:  $m_{\text{Cl}^-} = 0.1500 \text{ g} \times \frac{35.45}{143.32} = 0.0375 \text{ g}$

# Analysis Of Chloride Experiment 7 Lab Report

**Michael Brown**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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## **Table of Contents Analysis Of Chloride Experiment 7 Lab Report**

1. Understanding the eBook Analysis Of Chloride Experiment 7 Lab Report
  - The Rise of Digital Reading Analysis Of Chloride Experiment 7 Lab Report
  - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of Chloride Experiment 7 Lab Report
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Analysis Of Chloride Experiment 7 Lab Report
  - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of Chloride Experiment 7 Lab Report
  - Personalized Recommendations
  - Analysis Of Chloride Experiment 7 Lab Report User Reviews and Ratings
  - Analysis Of Chloride Experiment 7 Lab Report and Bestseller Lists
5. Accessing Analysis Of Chloride Experiment 7 Lab Report Free and Paid eBooks
  - Analysis Of Chloride Experiment 7 Lab Report Public Domain eBooks
  - Analysis Of Chloride Experiment 7 Lab Report eBook Subscription Services

- Analysis Of Chloride Experiment 7 Lab Report Budget-Friendly Options
- 6. Navigating Analysis Of Chloride Experiment 7 Lab Report eBook Formats
  - ePub, PDF, MOBI, and More
  - Analysis Of Chloride Experiment 7 Lab Report Compatibility with Devices
  - Analysis Of Chloride Experiment 7 Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Analysis Of Chloride Experiment 7 Lab Report
  - Highlighting and Note-Taking Analysis Of Chloride Experiment 7 Lab Report
  - Interactive Elements Analysis Of Chloride Experiment 7 Lab Report
- 8. Staying Engaged with Analysis Of Chloride Experiment 7 Lab Report
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Analysis Of Chloride Experiment 7 Lab Report
- 9. Balancing eBooks and Physical Books Analysis Of Chloride Experiment 7 Lab Report
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Analysis Of Chloride Experiment 7 Lab Report
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Analysis Of Chloride Experiment 7 Lab Report
  - Setting Reading Goals Analysis Of Chloride Experiment 7 Lab Report
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analysis Of Chloride Experiment 7 Lab Report
  - Fact-Checking eBook Content of Analysis Of Chloride Experiment 7 Lab Report
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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