

Chemistry 1C
Worksheet #3

1. At 2200°C, $K_p = 0.050$ for the reaction:



a. What is the partial pressure of NO in equilibrium with N_2 and O_2 that was placed in a flask at initial pressures of 0.80 and 0.20 atm?

	N_2	O_2	2NO
Initial	0.8	0.20	0
Change	-x	-x	+2x
Equilibrium	$0.8-x$	$0.20-x$	2x

$$K_p = (2x)^2 / ((0.80-x)(0.20-x))$$

Solve by quadratic formula and get $x = 0.039$ atm or $x = -0.052$. So bc you cannot have negative pressure, the answer is 0.039.

b. What happens to the concentration of the NO if we increase the temperature?

Increases

c. What happens to the concentration of O_2 when you add N_2 at constant volume? **Decreases**

d. What happens to the concentration of O_2 when you add Ar at constant volume? **No change**

2. A 1.00 L flask was filled with 2.0 mol $\text{SO}_2 (\text{g})$ and 2.0 mol $\text{NO}_2 (\text{g})$ and heated.

After equilibrium was reached, it was found that 1.3 moles NO_2 was present.

Assume that the reaction occurs under these conditions



a. Calculate the equilibrium constant, K_c , for the above reaction.

	SO_2	NO_2	SO_3	NO
Initial	2.0	2.0	0	0
Change	-x	-x	+x	+x
Equilibrium	$2.0-x$	$2.0-x$	+x	+x

$$K_c = \frac{[\text{SO}_3][\text{NO}]}{[\text{SO}_2][\text{NO}_2]} = \frac{x^2}{(2.0-x)^2}$$

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