AP Chemistry Chapter 4 Practice Exam

- The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

$$2H_2S + 3O_2 \rightarrow 2SO_2 + 2H_2O$$

 $SO_2 + 2H_2S \rightarrow 3S + 2H_2O$

How much sulfur (in grams) is produced from 21.5 grams of O₂? 21.5502 |mol 02 | 2503 | 35 | 32.0675 = 43.155

- A) 97.0 g
- B) 14.4 g
- C) 21.5 g
- none of these
- B
- Iron is produced from its ore by the reactions:

$$2C(e) + O_2(e) \rightarrow 2CO(e)$$

$$Fe_2O_3(s) + 3CO(g) \rightarrow 2Fe(s) + 3CO_3(g)$$

Closes:

How many moles of $O_2(g)$ are needed to produce 4.6 moles of Fe(s)?

- A) 2.3 mole O₂ B) 3.5 mole O-
- C) 4.6 mole O₂
- D) 6.9 mole O₂
- E) 13.8 mole O₂

- B
- Suppose the reaction Ca₂(PO₄)₂ + 3H₂SO₄ → 3CaSO₄ + 2H₂PO₄ is carried out starting with 153 g. of $Ca_2(PO_k)_2$

and 76.8 g of H_2SO_4 . How much phosphoric acid will be produced?

A) 76.7 gB) 51.1 g $153 g (a_3 Pdq)_2 | 1m_1 (a_4 Rd_2)_2 | 2 H_3 Pdq$ $1 Ca_4 Rd_2 | 1m_1 H_2 R$

- B) 51.1 g C) 229.8 g

 - D) 115.1 c E) 96.7 g
- 7 6.89 H2504 | Ind H2504 2 H2 POY 1 3 =

1

- - SO₂ reacts with H₂S as follows:

$$2H_2S + SO_2 \rightarrow 3S + 2H_2O$$

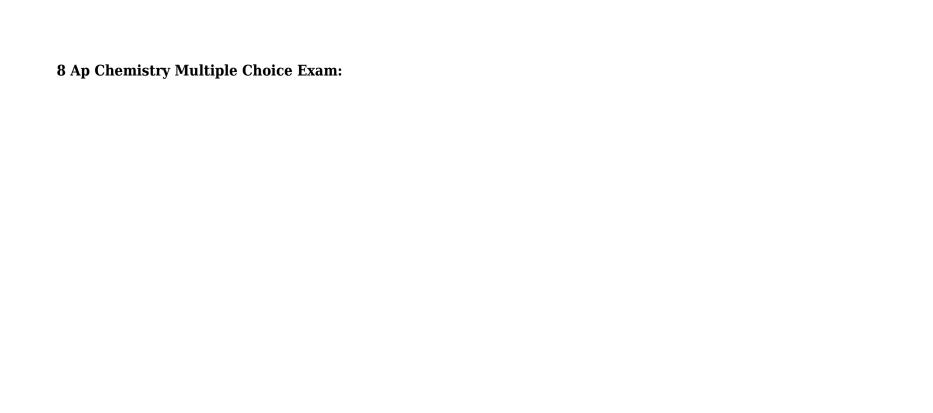
When 7.50 g of H₂S reacts with 12.75 g of SO₂, which statement applies?

- A) 6.38 g of sulfur are formed.
- B) 10.6 g of sulfur are formed.
- C) 0.0216 moles of H₂S remain.
- D) 1.13 g of H₂S remain.
- E) SO₂ is the limiting reagent.

4.6 mil Fersi 300 102 = 3.45mol02

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