

Chemistry

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(Chapter 1)(Solutions) (Intext Questions)

(Class 12)

Question 1.1:

Calculate the mass percentage of benzene (C_6H_6) and carbon tetrachloride (CCl_4) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride.

Answer 1.1:

$$\text{Mass percentage of } C_6H_6 = \frac{\text{Mass of } C_6H_6}{\text{Total mass of the solutions}} \times 100 \%$$

$$= \frac{\text{Mass of } C_6H_6}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{22}{22 + 122} \times 100 \% = 15.28 \%$$

$$\text{Mass percentage of } CCl_4 = \frac{\text{Mass of } CCl_4}{\text{Total mass of the solutions}} \times 100 \%$$

$$= \frac{\text{Mass of } CCl_4}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{122}{22 + 122} \times 100 \% = 84.72 \%$$

$$\text{Alternatively, Mass percentage of } CCl_4 = (100 - 15.28) \% = 84.72 \%$$

Question 1.2:

Calculate the mole fraction of benzene in solution containing 30% by mass in carbon tetrachloride.

Answer 1.2:

Let the total mass of the solution be 100 g and the mass of benzene be 30 g.

$$\therefore \text{Mass of carbon tetrachloride} = (100 - 30) \text{ g} = 70 \text{ g}$$

$$\text{Molar mass of benzene } (C_6H_6) = (6 \times 12 + 6 \times 1) \text{ g mol}^{-1} = 78 \text{ g mol}^{-1}$$

$$\therefore \text{Number of moles of } C_6H_6 = \frac{30}{78} \text{ mol} = 0.3846 \text{ mol}$$

$$\text{Molar mass of carbon tetrachloride } (CCl_4) = 1 \times 12 + 4 \times 35.5 \\ = 154 \text{ g mol}^{-1}$$

$$\therefore \text{Number of moles of } CCl_4 = \frac{70}{154} \text{ mol} = 0.4545 \text{ mol}$$

Thus, the mole fraction of C_6H_6 is given as:

$$= \frac{\text{Number of moles of } C_6H_6}{\text{Number of moles of } C_6H_6 + \text{Number of moles of } CCl_4}$$

$$= \frac{0.3846}{0.3846 + 0.4545}$$

$$= 0.458$$

Question 1.3:

Calculate the molarity of each of the following solutions:

(a) 30 g of $Co(NO_3)_2 \cdot 6H_2O$ in 4.3 L of solution

(b) 30 mL of 0.5 M H_2SO_4 diluted to 500 mL

Answer 1.3:

$$\text{Molarity is given by: } \text{Molarity} = \frac{\text{Moles of solute}}{\text{Volume of solution in litre}}$$

$$\text{(a) Molar mass of } Co(NO_3)_2 \cdot 6H_2O = 59 + 2(14 + 3 \times 16) + 6 \times 18 = 291 \text{ g mol}^{-1}$$

$$\therefore \text{Moles of } Co(NO_3)_2 \cdot 6H_2O = \frac{30}{291} \text{ mol} = 0.103 \text{ mol}$$

$$\text{Therefore, molarity} = \frac{0.103 \text{ mol}}{4.3 \text{ L}} = 0.023 \text{ M}$$

$$\text{(b) Number of moles present in 1000 mL of 0.5 M } H_2SO_4 = 0.5 \text{ mol}$$

$$\therefore \text{Number of moles present in 30 mL of 0.5 M } H_2SO_4 = \frac{0.5 \times 30}{1000} \text{ mol} = 0.015 \text{ mol}$$

$$\text{Therefore, molarity} = \frac{0.015 \text{ mol}}{0.5 \text{ L}} = 0.03 \text{ M}$$

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