

18.



According to the equation above, how many moles of potassium chlorate, KClO_3 , must be decomposed to generate 1.0 L of O_2 gas at standard temperature and pressure?

- A) $\frac{1}{3} \left(\frac{1}{22.4} \right)$ mol
 B) $\frac{1}{2} \left(\frac{1}{22.4} \right)$ mol
 C) $\frac{2}{3} \left(\frac{1}{22.4} \right)$ mol
 D) $\frac{3}{2} \left(\frac{1}{22.4} \right)$ mol
 E) $2 \left(\frac{1}{22.4} \right)$ mol

19. Which of the following ions has the same number of electrons as Br^- ?

- A) Ca^{2+}
 B) K^+
 C) Sr^{2+}
 D) I^-
 E) Cl^-

20. In solid methane, the forces between neighboring CH_4 molecules are best characterized as

- A) ionic bonds
 B) covalent bonds
 C) hydrogen bonds
 D) ion-dipole forces
 E) London (dispersion) forces

21. Of the following electron configurations of neutral atoms, which represents an atom in an excited state?

- A) $1s^2 2s^2 2p^2$
 B) $1s^2 2s^2 2p^2 3s^1$
 C) $1s^2 2s^2 2p^4 3s^1$
 D) $1s^2 2s^2 2p^6 3s^2 3p^2$
 E) $1s^2 2s^2 2p^6 3s^2 3p^1$

22. Which of the following is a nonpolar molecule that contains polar bonds?

- A) F_2
 B) CHCl_3
 C) CO_2
 D) HCl
 E) NH_3

23. The oxidation state that is common to aqueous ions of Fe, Mn, and Zn is

- A) +1
 B) +2
 C) +3
 D) +4
 E) +5

24. Which of the following shows the correct number of protons, neutrons, and electrons in a neutral cesium-134 atom?

	Protons	Neutrons	Electrons
A)	55	55	55
B)	55	79	55
C)	55	79	79
D)	79	55	79
E)	134	55	134

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