

1 A certain reaction is zero order in reactant A and second order in reactant B. What happens to the reaction rate when the concentrations of both reactants are doubled?

Choose one answer.

- a. The reaction rate remains the same. ✗
- b. The reaction increases by a factor of two. ✗
- c. The reaction increases by a factor of four. ✓
- d. The reaction increases by a factor of eight. ✗

2 A rate law relates which of the following?

Choose one answer.

- a. Reaction rate and temperature ✗
- b. Reaction rate and concentration ✓
- c. Temperature and concentration ✗
- d. Energy and concentration ✗

3 Determine the overall balanced equation for a reaction that has the following proposed mechanism. Which of the following is the acceptable rate law?
Step 1: $B_2 + B_2 \rightarrow E_3 + D$ slow
Step 2: $E_3 + A \rightarrow B_2 + C_2$ fast

Choose one answer.

- a. $B_2 + B_2 \rightarrow E_3 + D$ and $R = [B_2]^2$ ✗
- b. $A + B_2 \rightarrow C_2 + D$ and $R = [A][B_2]^2$ ✗
- c. $A + B_2 \rightarrow C_2 + D$ and $R = [B_2]^2$ ✓
- d. $A + B_2 \rightarrow C_2 + D$ and $R = [E_3][A]$ ✗

4 Fill in the blank. A species that changes the rate of a reaction but is neither consumed nor changed in that reaction is a(n) _____.

Choose one answer.

- a. Catalyst ✓
- b. Activated complex ✗
- c. Intermediate ✗
- d. Reactant ✗

5 If two molecules collide softly, what will most likely happen to the molecules?

Choose one answer.

- a. The molecules will be oriented favorably. ✗
- b. The molecules will not be oriented favorably. ✗
- c. The molecules are likely to react. ✗
- d. The molecules may rebound without reacting. ✓

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