

Researched expected results

Test 1

	Magnesium chloride	Calcium bromide	Strontium chloride	Barium chloride
Initial	Colourless solution	Colourless solution	Colourless solution	Colourless solution
After adding 10 drops of 0.5M NaOH solution	Slightly white precipitate	Slightly white precipitate	Slightly white precipitate	Colourless solution
After adding excess NaOH	White precipitate	Slightly white precipitate	Slightly white precipitate	Colourless solution
Reaction and Product formed	$\text{MgCl}_2 + 2\text{NaOH} \rightarrow \text{Mg(OH)}_2(\text{s}) + 2\text{NaCl}$	$\text{CaBr}_2 + 2\text{NaOH} \rightarrow \text{Ca(OH)}_2(\text{s}) + 2\text{NaBr}$	$\text{SrCl}_2 + 2\text{NaOH} \rightarrow \text{Sr(OH)}_2(\text{s}) + 2\text{NaCl}$	$\text{BaCl}_2 + \text{NaOH} \rightarrow \text{Ba(OH)}_2 + \text{NaCl}$

Test 2

	Magnesium chloride	Calcium bromide	Strontium chloride	Barium chloride
Initial	Colourless solution	Colourless solution	Colourless solution	Colourless solution
After adding 10 drops of 1.0M H ₂ SO ₄ solution	Slight white precipitate	Slight white precipitate	white precipitate	white precipitate
After adding excess H ₂ SO ₄	Colourless solution	Slight white precipitate	white precipitate	white precipitate
Reaction and Product formed	$\text{MgCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + 2\text{HCl}$	$\text{CaBr}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{SO}_2 + \text{Br}_2 + \text{CaSO}_4(\text{s})$	$\text{SrCl}_2 + 2\text{H}_2\text{SO}_4 \rightarrow \text{Sr(HSO}_4)_2(\text{s}) + 2\text{HCl}$	$\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4(\text{s}) + \text{HCl}$

Test 3

Observation of damp red litmus paper being held over the test tube containing a heated mixture of 10 drops of 0.1M ammonium chloride solution and 10 drops 0.5M sodium hydroxide solution.	Litmus paper turns blue
Reaction and product formed	$\text{NH}_4\text{Cl} + \text{NaOH} \rightarrow \text{NH}_3 + \text{NaCl} + \text{H}_2\text{O}$
Reason	Ammonium ions can be recognised in a solution. If ammonium ions are present when added to sodium hydroxide, they will be converted to ammonia gas and/or turn red litmus paper blue.

Test 4

Observation of red litmus paper with the fumes of 1cm ³ solution of 0.5M sodium hydroxide	Turns damp red litmus paper blue
Observations made after a few minutes of dampening red litmus paper with deionised water and placing it beside filter paper soaked in 1M ammonia solution inside a closed petri dish	Turns damp red litmus paper blue

Test 5

Solution	Volume (cm ³)	Molarity (M)	Observation
Limewater	1	0.02	Colourless solution
$\text{HCl} + \text{Na}_2\text{CO}_3 \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$	3	1.00	After adding hydrochloric acid and sodium carbonate to the limewater, the solution went from colourless to cloudy.

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Table of Contents A2 Chemistry 2013 Unofficial Mark Scheme

1. Understanding the eBook A2 Chemistry 2013 Unofficial Mark Scheme
 - The Rise of Digital Reading A2 Chemistry 2013 Unofficial Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying A2 Chemistry 2013 Unofficial Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an A2 Chemistry 2013 Unofficial Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from A2 Chemistry 2013 Unofficial Mark Scheme
 - Personalized Recommendations

- A2 Chemistry 2013 Unofficial Mark Scheme User Reviews and Ratings
- A2 Chemistry 2013 Unofficial Mark Scheme and Bestseller Lists
- 5. Accessing A2 Chemistry 2013 Unofficial Mark Scheme Free and Paid eBooks
 - A2 Chemistry 2013 Unofficial Mark Scheme Public Domain eBooks
 - A2 Chemistry 2013 Unofficial Mark Scheme eBook Subscription Services
 - A2 Chemistry 2013 Unofficial Mark Scheme Budget-Friendly Options
- 6. Navigating A2 Chemistry 2013 Unofficial Mark Scheme eBook Formats
 - ePub, PDF, MOBI, and More
 - A2 Chemistry 2013 Unofficial Mark Scheme Compatibility with Devices
 - A2 Chemistry 2013 Unofficial Mark Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of A2 Chemistry 2013 Unofficial Mark Scheme
 - Highlighting and Note-Taking A2 Chemistry 2013 Unofficial Mark Scheme
 - Interactive Elements A2 Chemistry 2013 Unofficial Mark Scheme
- 8. Staying Engaged with A2 Chemistry 2013 Unofficial Mark Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers A2 Chemistry 2013 Unofficial Mark Scheme
- 9. Balancing eBooks and Physical Books A2 Chemistry 2013 Unofficial Mark Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection A2 Chemistry 2013 Unofficial Mark Scheme
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine A2 Chemistry 2013 Unofficial Mark Scheme
 - Setting Reading Goals A2 Chemistry 2013 Unofficial Mark Scheme
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of A2 Chemistry 2013 Unofficial Mark Scheme
 - Fact-Checking eBook Content of A2 Chemistry 2013 Unofficial Mark Scheme

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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