Amended from AQA Chemistry Unit 1 Foundation Chemistry CHEM1 June 2010 (Question 1)

Fluorine forms many compounds that contain covalent bonds.

a i State the meaning of the term covalent bond.

Write an equation to show the formation of one molecule of CIFs from chlorine and fluorine molecules.

b Draw the shape of a dichlorodifluoromethane molecule (CCl₂F₂) and the shape of a chlorine trifluoride molecule (ClF₃), include any lone pairs of electrons that influence the shape.
Shape of CCl₂F₂
Shape of CIF₃

(2 marks)

 Suggest the strongest type of intermolecular force between CCl₂F₂ molecules.

- d BF₃ is a covalent molecule that reacts with an F⁻ ion to form a BF₄-ion.
 - i Name the type of bond formed when a molecule of BF₃ reacts with an F⁻ ion. Explain how this bond is formed.

Type of bond

Explanation

(2 marks)

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John Atkinson, Carol Hibbert

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