

Name: _____

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ALGEBRA 2: FINAL EXAM REVIEW 2012-2013 1ST SEMESTER

1. Simplify the expression:

$$= (-2x - 3x^2 + 7) + (-5x^2 + 7 + 3x)$$

$$= \frac{8x - 56x^2}{8x}$$

$$= \frac{4x^2y^3 - 4x^2y - 32x^2y^2}{4x^2y^2}$$

$$= -5x^2y^2(4x - 2x^2y^2 + y)$$

2. Multiply two binomials:

$$= (4x - 1)(3x - 7)$$

$$= (3x - 2y)^2$$

$$= 8x(2x - 1)^2$$

3. Polynomial Long Division:

$$= \text{Divide } 3x^3 + x^2 + 9x^2 + 2x - 3 \text{ by } x^2 + 2$$

4. Factor the following polynomials:

$$= x^2 + 7x + 6$$

$$= 2x^2 + 7x + 3$$

$$= x^2 - 27$$

$$5. \text{ Simplify: } \frac{x^2 - 64}{x^2 - 8x + 8}$$

$$6. \text{ In lowest terms, the quotient of } \frac{x^2 + 8x + 16}{x - 3} \div \frac{x^2 - 16}{x + 3} \text{ is}$$

7. Adding & Subtracting Rational Expressions:

$$= \frac{6x}{y^2z^2} + \frac{7}{y^2z^2}$$

$$= \frac{5}{x^2 - 10x + 16} - \frac{6}{x - 2}$$

Algebra Second Semester Final Answers

Michal Rosen-Zvi



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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
 one mole of HCl would be fully neutralized by one mole of

naoh naoh

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide naoh and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{NaOH} \cdot \text{Na}_2\text{C}_2\text{O}_4 \cdot 4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web this point of titration is referred to as the equivalence point considering the 1:1 stoichiometry of this acid base reaction
 $\text{NaOH (aq)} + \text{C}_6\text{H}_4\text{COOH (aq)} \rightarrow \text{C}_6\text{H}_4\text{COONa (aq)} + \text{H}_2\text{O (l)}$ the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the