

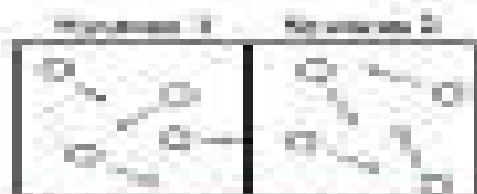
Unit 5: Thermochemistry**5.1: Heat Transfer and Thermal Equilibrium****5.2: Heat Capacity and Calorimetry****5.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles that move fast enough, the more kinetic energy we will see. Temperature is a scalar. The Boltzmann-probability curves show that the average velocity we will see for the distribution is increasing from $T1$ to $T3$, so the average temperature is increasing.

3. Sketch what you expect to represent the relative velocities of the particles in the following two gases, samples of the same gas.



- a. In System 1, I see a higher or lower temperature than System 2? Justify your answer.
 System 1 is at a higher temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more spread out than in System 2, so the Boltzmann-probability curve for System 1 is higher and broader than the curve for System 2. The area under the curve for System 1 is shaded pink, and the area under the curve for System 2 is shaded light blue. The peak of the curve for System 1 is at a higher energy than the peak of the curve for System 2, so the average velocity of the particles in System 1 is higher than the average velocity of the particles in System 2. The Boltzmann-probability curve for System 1 is higher and broader than the curve for System 2, so the average velocity of the particles in System 1 is higher than the average velocity of the particles in System 2.
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.
 Thermal energy will be transferred from System 1 to System 2. The particles in System 1 are moving faster than the particles in System 2, so the average velocity of the particles in System 1 is higher than the average velocity of the particles in System 2. The Boltzmann-probability curve for System 1 is higher and broader than the curve for System 2, so the average velocity of the particles in System 1 is higher than the average velocity of the particles in System 2. The area under the curve for System 1 is shaded pink, and the area under the curve for System 2 is shaded light blue. The peak of the curve for System 1 is at a higher energy than the peak of the curve for System 2, so the average velocity of the particles in System 1 is higher than the average velocity of the particles in System 2.

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Steven S. Zumdahl



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